# d<sup>2</sup>sp<sup>3</sup> Hybridization

- d<sup>2</sup>sp<sup>3</sup> hybridization enables an octahedral arrangement for six pairs of electrons surrounding a central atom.
  - Forms from two d orbitals, one s orbital, and three p orbitals.
- Examples
  - OSF6
  - XeF<sub>4</sub> (Xe has two lone pairs)



# **Evan P Silberstein 20answers Atomic Structure**

**Christian G. Meyer** 

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