

## Quiz 4 – Potential Energy Diagrams

1. The addition of a catalyst to a reaction provides an alternative mechanism with

- A. Lower activation energy and lower reaction rate
- B. Lower activation energy and higher reaction rate**
- C. Higher activation energy and lower reaction rate
- D. Higher activation energy and higher reaction rate

2. Consider the following Reaction



The correct equation including the heat term is

- A.  $\text{N}_{2(g)} + \text{O}_{2(g)} + 90 \text{ kJ} \longrightarrow 2\text{NO}_{(g)}$
- B.  $\text{N}_{2(g)} + \text{O}_{2(g)} + 180 \text{ kJ} \longrightarrow 2\text{NO}_{(g)}$**
- C.  $\text{N}_{2(g)} + \text{O}_{2(g)} \longrightarrow 2\text{NO}_{(g)} + 90 \text{ kJ}$
- D.  $\text{N}_{2(g)} + \text{O}_{2(g)} \longrightarrow 2\text{NO}_{(g)} + 180 \text{ kJ}$

3. A forward reaction has an activation energy of 50 kJ and a  $\Delta H$  of  $-100 \text{ kJ}$ .  
The PE diagram, which describes this reaction, is

A.

4. Consider the following potential energy diagram

The Activation energy for the forward reaction is

- A. 25 kJ
- B. 50 kJ
- C. 75 kJ**
- D. 125 kJ

5. Consider the following reaction



# Multiple Choice Questions Energy Diagrams

**Arshad Iqbal**



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### **Table of Contents Multiple Choice Questions Energy Diagrams**

1. Understanding the eBook Multiple Choice Questions Energy Diagrams
  - The Rise of Digital Reading Multiple Choice Questions Energy Diagrams
  - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Questions Energy Diagrams
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Multiple Choice Questions Energy Diagrams
  - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Questions Energy Diagrams
  - Personalized Recommendations
  - Multiple Choice Questions Energy Diagrams User Reviews and Ratings
  - Multiple Choice Questions Energy Diagrams and Bestseller Lists

5. Accessing Multiple Choice Questions Energy Diagrams Free and Paid eBooks
  - Multiple Choice Questions Energy Diagrams Public Domain eBooks
  - Multiple Choice Questions Energy Diagrams eBook Subscription Services
  - Multiple Choice Questions Energy Diagrams Budget-Friendly Options
6. Navigating Multiple Choice Questions Energy Diagrams eBook Formats
  - ePub, PDF, MOBI, and More
  - Multiple Choice Questions Energy Diagrams Compatibility with Devices
  - Multiple Choice Questions Energy Diagrams Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Multiple Choice Questions Energy Diagrams
  - Highlighting and Note-Taking Multiple Choice Questions Energy Diagrams
  - Interactive Elements Multiple Choice Questions Energy Diagrams
8. Staying Engaged with Multiple Choice Questions Energy Diagrams
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Multiple Choice Questions Energy Diagrams
9. Balancing eBooks and Physical Books Multiple Choice Questions Energy Diagrams
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Multiple Choice Questions Energy Diagrams
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Multiple Choice Questions Energy Diagrams
  - Setting Reading Goals Multiple Choice Questions Energy Diagrams
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Multiple Choice Questions Energy Diagrams
  - Fact-Checking eBook Content of Multiple Choice Questions Energy Diagrams
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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