

Name : _____ Date : _____

Acids and Bases Strength

Write True (T) or False (F).

- i. Stronger acids have weaker conjugate bases and weaker acids have stronger conjugate bases. _____
- ii. Stronger bases have stronger conjugate acids and weaker bases have weaker conjugate acids. _____

Write equations that show how the following reactions occur in water.

- i. F^- acts as a base. _____
- ii. HNO_2 acts as an acid. _____
- iii. $\text{Fe}(\text{H}_2\text{O})_6^{3+}$ acts as an acid. _____
- iv. HCO_3^- acts as a base. _____
- v. HCO_3^- acts as an acid. _____
- vi. $\text{Al}(\text{H}_2\text{O})_5(\text{OH})^{2+}$ acts as a base. _____

What volume of 0.45 M LiOH would be needed to neutralize 60.0 mL of 0.15 M HI?

Calculate the pH values of the following solutions of strong acids and bases.

- i. 0.001 M HCl
- ii. 0.76 M KOH
- iii. 2.8×10^{-4} M $\text{Ba}(\text{OH})_2$

Lecture 29 Acids And Bases I Worksheet

Brendan G. Carr



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