

Question Number	Acceptable Answers	Reject	Mark
22(b)(iii)	$\Delta H_{\text{reaction}} = + \frac{2.3(10)}{0.2(00)} \times 2 = +23.1 \text{ (kJ mol}^{-1}\text{)}$ $= +23 \text{ (kJ mol}^{-1}\text{) to 2 sf}$ First mark – correct computation of $\Delta H_{\text{reaction}}$: 2 x [answer to (b)(i) in kJ ÷ answer to (b)(ii) in mol] (1) Second mark – stand alone, for correct rounding: A final answer to two sf (1) Third mark – stand alone, for giving a + sign for endothermic reaction: + sign in front of final answer (1) NOTE: + 12 (kJ mol⁻¹) scores (2) (i.e. the 2nd and 3rd marks)	Incorrect units given by the candidate (no 3 rd scoring point)	3

Question Number	Acceptable Answers	Reject	Mark
22(c)(i)	(Average amount of) energy/enthalpy required to break one mole of (covalent) bonds ALLOW Energy change/enthalpy change to break one mole of (covalent) bonds (1) (in the) gas / gaseous (state) (1)	Energy/enthalpy released OR 'Bonds formed/made' OR 1 mol of compound for 1st mark	2

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