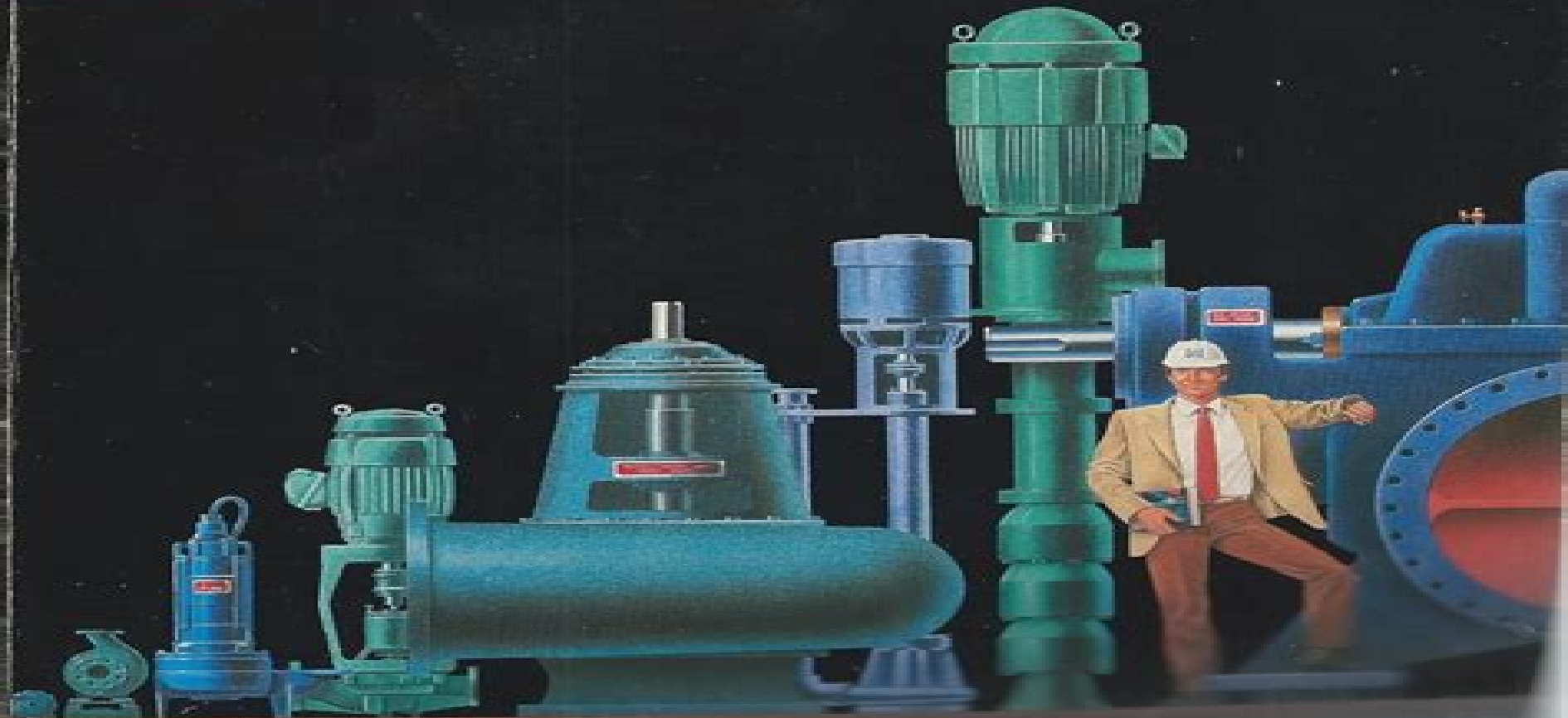


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 what do the following terms mean oxidation oxidation is a loss of electrons reduction reduction is a gain of electrons what is
 a helpful way to remember this

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 oxidation $\text{Co} \rightarrow \text{Co}^{2+} + 2\text{e}^-$ reduction $\text{Fe}^{2+} + 2\text{e}^- \rightarrow \text{Fe}$ 24 3 $\text{Ag} + \text{Ni} \rightarrow \text{Ni}^{3+} + 3\text{Ag}$ oxidation reduction 25 $\text{Cu}^{2+} + \text{Pb} \rightarrow \text{Pb}^{2+} + \text{Cu}$ oxidation reduction 26 O_2

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 magnitude of change then identify the element oxidized and reduced as well as the oxidizing agent and reducing agent a I_2O_5
 $5\text{CO} + \text{I}_2\text{O}_5 \rightarrow 5\text{CO}_2 + \text{I}_2$ 5 $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$ 5 $\text{C} + \text{O}_2 \rightarrow \text{CO}_2$

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