

- Instructions: (1) Substantiate your answers with formulas for finding using one or more of the following methods:
 (2) Balance each equation.
 (3) Identify the type of chemical equation, decomposition, single replacement, double replacement, or combination.

1. Aluminum metal reacts with oxygen gas to form aluminum oxide and oxygen.
 aluminum + oxygen →

Type: _____

balanced: $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$

2. Iron(III) sulfate reacts with barium hydroxide (aq) solution.

Type: _____

balanced: $3\text{Fe}_2(\text{SO}_4)_3 + 3\text{Ba}(\text{OH})_2 \rightarrow 6\text{Fe}(\text{OH})_3 + 9\text{BaSO}_4$

3. Calcium phosphate reacts with phosphoric acid to produce calcium dihydrogen phosphate and phosphoric acid. Write balanced chemical equation for this reaction.

Type: _____

balanced: $\text{Ca}_3(\text{PO}_4)_2 + 4\text{H}_3\text{PO}_4 \rightarrow 3\text{CaH}_2\text{P}_2\text{O}_7 + 2\text{H}_2\text{O}$

4. Ammonium phosphate reacts to produce nitrogen gas, phosphoric acid, and water.

Type: _____

balanced: $(\text{NH}_4)_3\text{PO}_4 \rightarrow \text{N}_2 + \text{H}_2\text{O} + \text{H}_3\text{PO}_4$

5. Ammonium phosphate reacts with calcium hydroxide (aq) solution, to produce calcium phosphate (s) and water (l).

Type: _____

balanced: $(\text{NH}_4)_3\text{PO}_4 + 3\text{Ca}(\text{OH})_2 \rightarrow \text{Ca}_3(\text{PO}_4)_2 + 6\text{H}_2\text{O}$

Even More Fun With Equations 6 33 Answers

M Carnoy



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