Q6 When water was added to the anhydrate of CuSO₄(x), did the reaction proceed in the direction of the reactants or products?

B.2 Reaction of Copper(II) Ion and Hydroxide

	Test tube 1	Test tube 2
Initial color of CuCl ₂ solution (1)	reansparent	transparent
	cloudes blue	cloudy bive
Change in appearance (3)	more cloudy solution	The second secon
Change in appearance (4)		cicar lighter tive

- Q7 In test tube 1, did the addition of OH⁻ cause the reaction to go in the direction of the reactants or products? Explain.
- Q8 In test tube 2, did the addition of H₃O⁺ cause the reaction to go in the direction of the reactants or products? Explain.

C. Changing Equilibrium Conditions: Le Châtelier's Principle

C.1 Effect of Changing Reactant Concentration

Test tube	Color	Is color deeper or lighter red compared to the reference?	[FeSCN ²⁺] increases or decreases?
1. Reference	blood red		
2. Fe ³⁺ added	danc blood	deeper	
3. SCN ⁻ added	danc blood	deeper	
4. Fe ³⁺ removed	light blood	lighter	

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