

Key

Ideal Gas Law Worksheet $PV = nRT$

Use the ideal gas law, " $PerV=nRT$ ", and the universal gas constant $R = 0.0821 \frac{L \cdot atm}{K \cdot mol}$ to solve the following problems:

If pressure is needed in kPa then convert by multiplying by $101.3 kPa / 1 atm$ to get
 $R = 8.31 kPa \cdot L / (K \cdot mole)$

- 1) If I have 4 moles of a gas at a pressure of 5.6 atm and a volume of 12 liters, what is the temperature?

$$PV = nRT$$

$$T = \frac{PV}{nR} = \frac{(5.6 atm)(12 L)}{4 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol}}$$

$$T = 204.63 K$$

- 2) If I have an unknown quantity of gas at a pressure of 1.2 atm, a volume of 31 liters, and a temperature of $87^{\circ}C$, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{1.2 atm \cdot 31 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 360 K}$$

$$n = 1.2586 mol$$

- 3) If I contain 3 moles of gas in a container with a volume of 60 liters and at a temperature of 400 K, what is the pressure inside the container?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{3 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 400 K}{60 L}$$

$$P = 1.642 atm$$

$$or \quad P = 166.29 kPa$$

- 4) If I have 7.7 moles of gas at a pressure of 0.09 atm and at a temperature of $56^{\circ}C$, what is the volume of the container that the gas is in?

$$PV = nRT$$

$$V = \frac{nRT}{P} = \frac{7.7 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 327 K}{0.09 atm}$$

$$V = 2310.93 L$$

- 5) If I have 17 moles of gas at a temperature of $67^{\circ}C$, and a volume of 88.89 liters, what is the pressure of the gas?

$$PV = nRT$$

$$P = \frac{nRT}{V} = \frac{17 mol \cdot 0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 390 K}{88.89 L}$$

$$P = 5.34 atm$$

$$or \quad P = 540.61 kPa$$

- 6) If I have an unknown quantity of gas at a pressure of 0.5 atm, a volume of 25 liters, and a temperature of 300 K, how many moles of gas do I have?

$$PV = nRT$$

$$n = \frac{PV}{RT} = \frac{0.5 atm \cdot 25 L}{0.0821 \frac{L \cdot atm}{K \cdot mol} \cdot 300 K}$$

$$n = 0.5075 mol$$

Ideal Gas Law Problems Answer Key

Abhay Kumar Singh

A red circular graphic with a gradient, appearing as a semi-circle or a partial circle, located to the right of the author's name.

Ideal Gas Law Problems Answer Key:

The book delves into Ideal Gas Law Problems Answer Key. Ideal Gas Law Problems Answer Key is an essential topic that needs to be grasped by everyone, ranging from students and scholars to the general public. The book will furnish comprehensive and in-depth insights into Ideal Gas Law Problems Answer Key, encompassing both the fundamentals and more intricate discussions.

1. This book is structured into several chapters, namely:
 - Chapter 1: Introduction to Ideal Gas Law Problems Answer Key
 - Chapter 2: Essential Elements of Ideal Gas Law Problems Answer Key
 - Chapter 3: Ideal Gas Law Problems Answer Key in Everyday Life
 - Chapter 4: Ideal Gas Law Problems Answer Key in Specific Contexts
 - Chapter 5: Conclusion
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4. In chapter 3, this book will examine the practical applications of Ideal Gas Law Problems Answer Key in daily life. This chapter will showcase real-world examples of how Ideal Gas Law Problems Answer Key can be effectively utilized in everyday scenarios.
5. In chapter 4, the author will scrutinize the relevance of Ideal Gas Law Problems Answer Key in specific contexts. The fourth chapter will explore how Ideal Gas Law Problems Answer Key is applied in specialized fields, such as education, business, and technology.
6. In chapter 5, this book will draw a conclusion about Ideal Gas Law Problems Answer Key. The final chapter will summarize the key points that have been discussed throughout the book.

The book is crafted in an easy-to-understand language and is complemented by engaging illustrations. This book is highly recommended for anyone seeking to gain a comprehensive understanding of Ideal Gas Law Problems Answer Key.

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