

Ideal Gas Law Practice Worksheet

Solve the following problems using the ideal gas law:

- 1) How many moles of gas does it take to occupy 120 liters at a pressure of 2.3 atmospheres and a temperature of 340 K? $PV = nRT$
 $V = 120\text{ L}$
 $P = 2.3\text{ atm}$
 $T = 340\text{ K}$
 $n = ?$

$$(2.3\text{ atm})(120\text{ L}) = n(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(340\text{ K})$$

$$\boxed{n = 9.9\text{ mol}}$$
- 2) If I have a 50 liter container that holds 45 moles of gas at a temperature of 200° C, what is the pressure inside the container?
 $V = 50\text{ L}$
 $n = 45\text{ mol}$
 $T = 200^{\circ}\text{C} = 473\text{ K}$
 $P = ?$

$$(P)(50\text{ L}) = (45\text{ mol})(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(473\text{ K})$$

$$\boxed{P = 30\text{ atm}} \quad \text{or} \quad \boxed{4000\text{ kPa}}$$
- 3) It is not safe to put aerosol canisters in a campfire, because the pressure inside the canisters gets very high and they can explode. If I have a 1.0 liter canister that holds 2 moles of gas, and the campfire temperature is 1400° C, what is the pressure inside the canister?
 $V = 1.0\text{ L}$
 $n = 2\text{ mol}$
 $T = 1400^{\circ}\text{C} = 1673\text{ K}$
 $P = ?$

$$(P)(1.0\text{ L}) = (2\text{ mol})(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(1673\text{ K})$$

$$\boxed{P = 275\text{ atm} = 280\text{ atm}}$$

$$\text{or } 28,000\text{ kPa}$$
- 4) How many moles of gas are in a 30 liter scuba canister if the temperature of the canister is 300 K and the pressure is 200 atmospheres?
 $V = 30\text{ L}$
 $T = 300\text{ K}$
 $P = 200\text{ atm}$
 $n = ?$

$$(200\text{ atm})(30\text{ L}) = n(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(300\text{ K})$$

$$\boxed{n = 243.6\text{ mol} = 240\text{ mol}}$$
- 5) I have a balloon that can hold 100 liters of air. If I blow up this balloon with 3.0 moles of oxygen gas at a pressure of 1.0 atmosphere, what is the temperature of the balloon?
 $V = 100\text{ L}$
 $n = 3\text{ mol}$
 $P = 1\text{ atm}$
 $T = ?$

$$(1\text{ atm})(100\text{ L}) = (3\text{ mol})(0.0821 \frac{\text{L}\cdot\text{atm}}{\text{K}\cdot\text{mol}})(T)$$

$$\boxed{T = 406\text{ K} = 133^{\circ}\text{C} = 130^{\circ}\text{C}}$$

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Holt Chemistry Ralph Thomas Myers, 2004 *Merrill Chemistry* Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 **A Guide to Teaching in the Active Learning Classroom** Paul Baepler, J. D. Walker, D. Christopher Brooks, Kem Saichaie, Christina I. Petersen, 2023-07-03 While Active Learning Classrooms or ALCs offer rich new environments for learning they present many new challenges to faculty because among other things they eliminate the room's central focal point and disrupt the conventional seating plan to which faculty and students have become accustomed. The importance of learning how to use these classrooms well and to capitalize on their special features is paramount. The potential they represent can be realized only when they facilitate improved learning outcomes and engage students in the learning process in a manner different from traditional classrooms and lecture halls. This book provides an introduction to ALCs briefly covering their history and then synthesizing the research on these spaces to provide faculty with empirically based practical guidance on how to use these unfamiliar spaces effectively. Among the questions this book addresses are: How can instructors mitigate the apparent lack of a central focal point in the space? What types of learning activities work well in the ALCs and take advantage of the affordances of the room? How can teachers address familiar classroom management challenges in these unfamiliar spaces? If assessment and rapid feedback are critical in active learning, how do they work in a room filled with circular tables and no central focus point? How do instructors balance group learning with the needs of the larger class? How can students be held accountable when many will necessarily have their backs facing the instructor? How can instructors evaluate the effectiveness of their teaching in these spaces? This book is intended for faculty preparing to teach in or already working in this new classroom environment for administrators planning to create ALCs or experimenting with provisionally designed rooms and for faculty developers helping teachers transition to using these new spaces.

American Journal of Physics, 2005 *Simulations and Student Learning* Matthew Schnurr, Anna MacLeod, 2021-01-04 The book underlines the value of simulation based education as an approach that fosters authentic engagement and deep learning. **Spreadsheets for Chemists** Gordon Filby, 1995 A practical guide. *Spreadsheets for Chemists* shows chemists of all levels how to use spreadsheet programs in their daily work. It highlights the possibilities provided by Lotus 1 2 3, the most widely used spreadsheet program in the sciences. Apart from hundreds of example fragments, it features detailed discussion of the most relevant functions and all the macro commands. An accompanying diskette containing 57 worksheets involving many different fields of chemical research and teaching. An extensive glossary of spreadsheet terms. Three appendices covering 1 2 3's competitors and add-in packages, the use of Windows based spreadsheets, and how what-if analysis and back-solving is applied. Although the disk examples were developed for Lotus 1 2 3 DOS Versions 2 x x 2 4, the worksheets are compatible with the newer Windows versions and those of their main competitors Borland's Quattro Pro and Microsoft's Excel. Several compatible spreadsheets, AsEasyAs Pro, Quake, might also be used as inexpensive alternatives. The author is a

senior scientist at the Nuclear Research Centre in Karlsruhe Germany He has been using spreadsheet software for nearly ten years successfully in a variety of chemical problems

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The Software Encyclopedia, 1985 **Ideal Gas Law 51 Success Secrets - 51 Most Asked Questions on Ideal Gas Law - What You Need to Know** Philip Hicks, 2014-10-07 It's a brand new Ideal gas law world There has never been a Ideal gas law Guide like this It contains 51 answers much more than you can imagine comprehensive answers and extensive details and references with insights that have never before been offered in print Get the information you need fast This all embracing guide offers a thorough view of key knowledge and detailed insight This Guide introduces what you want to know about Ideal gas law A quick look inside of some of the subjects covered Atmospheric thermodynamics Overview Thermodynamic instruments Thermodynamic meters Glossary of engineering I Idealization Limits on use Perfect gas Stoichiometry Water vapor Water vapor and dry air density calculations at 0 C Equipartition theorem Perfection Physics and chemistry Glossary of chemistry terms U Fusion energy 1960s Timeline of low temperature technology 19th century Gas Avogadro's law Hot air balloon List

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