

Graphing Exercise #1

Richard C. Weast, ed., "CRC Handbook of Physics and Chemistry",
Chemical Section, 56th Edition.

Equation (22.30.1) is:

$$P/P^0 = \exp(-u/kT)$$

For this exercise for this subject and student, we made a few assumptions about the behavior of the gas. We called this our Kinetic Molecular Theory model (KMT).

1. The gas is a perfect gas (ideal gas).
2. The gas is a perfect gas (ideal gas) and is in equilibrium with the liquid phase.
3. The gas is a perfect gas (ideal gas) and is in equilibrium with the liquid phase.

Notice that these three assumptions are not identical. There are three different assumptions that we made for the Kinetic Molecular Theory. The Kinetic Molecular Equation is the gas equation is:

$$\left[P - \left(\frac{a}{V^2} \right) \right] (V - b) = nRT$$

The two terms that are subtracted are a/V^2 and b . The term a/V^2 is the attractive force and b is the volume of the gas.

We assume that the Kinetic Molecular Theory is a perfect gas (ideal gas) and is in equilibrium with the liquid phase. We assume that the Kinetic Molecular Theory is a perfect gas (ideal gas) and is in equilibrium with the liquid phase.

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Ideal Gas Graphing Exercise Answers

JR Anderson



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