

Answer following



- 1. Substrate and temperature:  $20^{\circ}\text{C}$   $30^{\circ}\text{C}$   $40^{\circ}\text{C}$
- 2. A graph of initial velocity:  $100\text{ mg/min}$   $200\text{ mg/min}$   $300\text{ mg/min}$
- 3. A graph of initial velocity:  $100\text{ mg/min}$   $200\text{ mg/min}$   $300\text{ mg/min}$

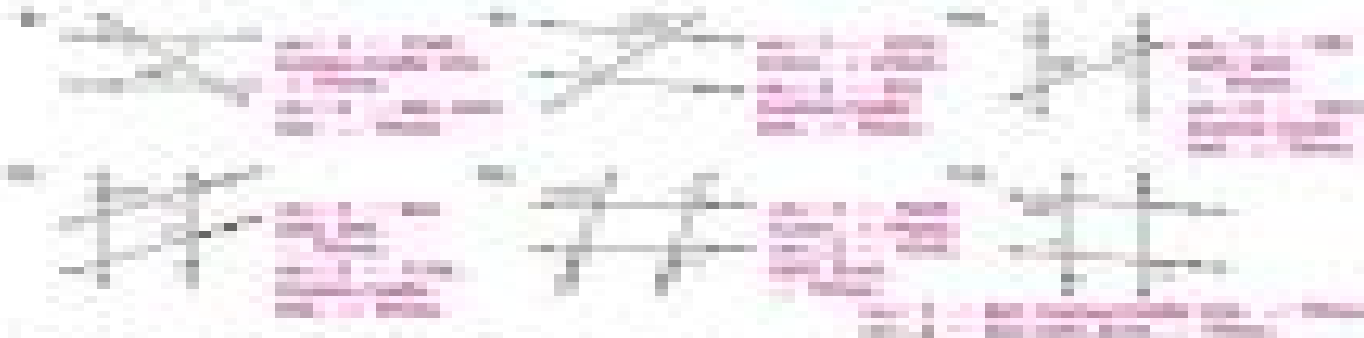
Calculate all values of  $K_m$  and  $V_{max}$  for the following reaction. Assume Michaelis-Menten kinetics and that the reaction is reversible.

- 1.  $K_m = 100\text{ mg/min}$   $V_{max} = 300\text{ mg/min}$
- 2.  $K_m = 100\text{ mg/min}$   $V_{max} = 300\text{ mg/min}$
- 3.  $K_m = 100\text{ mg/min}$   $V_{max} = 300\text{ mg/min}$
- 4.  $K_m = 100\text{ mg/min}$   $V_{max} = 300\text{ mg/min}$



Answer 100

Graph the following data. Assume the reaction is reversible and that the reaction is reversible.



Graph the following data. Assume the reaction is reversible and that the reaction is reversible.



Graph the following data. Assume the reaction is reversible and that the reaction is reversible.

# Geometry Lesson 10 6 Prentice Hall Answers

**Douglas Brooks**



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