Learning Target: Use isotope notation to determine: element name/symbol, atomic number, number of electrons, number of neutrons, number of protons, mass number, atomic number, atomic mass. Isotope Notation:

- 1. Here are three isotopes of an element: 6^{12} C 6^{13} C 6^{14} C
 - a. The element is: ____earbon_____
 - b. The number 6 refers to the ____atomic number____
 - c. The numbers 12, 13, and 14 refer to the ____atomic mass_
 d. How many protons and neutrons are in the first isotope? P=6 N=6
 - e. How many protons and neutrons are in the second isotope? P=6 N=7
 - f. How many protons and neutrons are in the third isotope? __P=6 N=7____
- 2. Complete the following chart:

Isotope name	atomic#	mass #	# of protons	# of neutrons	# of electrons	Isotope Notation
uranium-235	92	235	92	143	92	235 92-U
uranium-238	92	238	92	146	92	238 92- U
boron-10	5	10	5	5	5	*** **B
boron-11	5	11	5	6	5	sB

3.

Element	Symbol	Atomic Number	Number of electrons	Number of Neutrons	Mass number	Isotope Name	Isotope Notation
Helium	He	2	2	2	4	Helium-4	₄ ₂He
Titanium	Ti	22	22	28	50	Titanium-50	50 22 T/
Tantalum	Ta	73	73	108	181	Tantalum- 181	23 Ta
Gallium	Ga	31	31	39	70	Gallium-70	³⁰ 31 Ga
Carbon	C	6	6	7	13	Carbon-13	13C
Radium	Ra	88	88	218	226	Radium-226	seRa
Bismuth	Bi	83	83	127	210	Bismuth-210	216 83 B i

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