

Q	Part	Sub Part	Marking Guidance	Mark	Comments
5	(a)		<p>W is CuCl_4^{2-}</p> <p>Yellow-green/yellow/green</p> <p>$[\text{Cu}(\text{H}_2\text{O})_6]^{2+} + 4\text{Cl}^- \rightarrow \text{CuCl}_4^{2-} + 6\text{H}_2\text{O}$</p>	<p>1</p> <p>1</p> <p>1</p>	<p>Not necessary to indicate solution</p> <p>Do not allow precipitate/solid</p> <p>Allow $+ 4\text{HCl} \rightarrow 4\text{H}^+$</p>
5	(b)		<p>X is $\text{Cu}(\text{H}_2\text{O})_4(\text{OH})_2$</p> <p>Blue precipitate/solid</p> <p>$[\text{Cu}(\text{H}_2\text{O})_6]^{2+} + 2\text{NH}_3 \rightarrow \text{Cu}(\text{H}_2\text{O})_4(\text{OH})_2 + 2\text{NH}_4^+$</p>	<p>1</p> <p>1</p> <p>1</p>	<p>Allow $\text{Cu}(\text{OH})_2$/copper hydroxide</p> <p>Ignore shades</p> <p>Allow any balanced equation/equations leading to this hydroxide or $\text{Cu}(\text{OH})_2$ But must use ammonia</p>
5	(c)		<p>Y is $[\text{Cu}(\text{NH}_3)_4(\text{H}_2\text{O})_2]^{2+}$</p> <p>Deep/dark/royal <u>blue solution</u></p> <p>$\text{Cu}(\text{H}_2\text{O})_4(\text{OH})_2 + 4\text{NH}_3 \rightarrow [\text{Cu}(\text{NH}_3)_4(\text{H}_2\text{O})_2]^{2+} + 2\text{H}_2\text{O} + 2\text{OH}^-$</p>	<p>1</p> <p>1</p> <p>1</p>	<p>QoL</p> <p>Accept equation for formation from $\text{Cu}(\text{OH})_2$</p>
5	(d)		<p>Z is CuCO_3</p> <p>Green solid/precipitate</p> <p>$[\text{Cu}(\text{H}_2\text{O})_6]^{2+} + \text{CO}_3^{2-} \rightarrow \text{CuCO}_3 + 6\text{H}_2\text{O}$</p>	<p>1</p> <p>1</p> <p>1</p>	<p>Allow copper carbonate</p> <p>Allow blue-green precipitate</p>
5	(e)	(i)	<p>$\text{Cu}^{2+}(\text{aq}) + \text{Fe}(\text{s}) \rightarrow \text{Cu}(\text{s}) + \text{Fe}^{2+}(\text{aq})$</p> <p>Blue</p> <p>Green</p>	<p>1</p> <p>1</p> <p>1</p>	<p>Allow hydrated ions</p> <p>State symbols not essential but penalise if wrong</p> <p>Do not allow description of solids</p> <p>Allow yellow/(red-)brown/orange</p>

Isa Paper Carbonate Ore Aqa

RD Boyd



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