

Pre-Ap Chemistry Unit 6 – review for finals

These are the topics you should know and be able to answer questions about:

1. Chemical reactions

- a. What are some examples of evidence of chemical reactions?
 - i. Evidence that a chemical change *may have occurred* can include:
 - ii. Production of bubbles (gas)
 - iii. Formation of a precipitate (solid)
 - iv. Production of heat and light
 - v. Color change
 - vi. Noise
 - vii. Ultimately, only the formation of a NEW substance can be absolute proof that a chemical change has occurred.
- b. What is the difference between chemical and physical reactions?
 - i. Chemical reactions produce new products – what you start with and end with are different
 - ii. Physical reactions change the appearance of the substance, but you still have the same substance that you start with

2. Balance equations.

- a. Be able to balance equations
 - i. Either use math or tiddlywinks or draw a picture
- b. Why do we have to balance equations
 - i. Due to the law of conservation of mass. Mass is not created or destroyed during a chemical reaction. Due to this, the mass of the reactants must equal the mass of the products. Which means that there must be a balanced mass on each side of the reaction
- c. What are we trying to balance when we balance equations?
 - i. We need the same number of atoms of each element on each side of the arrow
- d. If you have balanced an equation and notice that the coefficients can be reduced, what do you? And why?
 - i. Reduce the coefficients.
 - ii. We need the lowest ratio of coefficients
- e. Draw for me the particle representation of a balanced equation (circles) for the formation of water ($\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$)



- f. In order to balance the equation $\text{KClO}_3 \rightarrow \text{KCl} + \text{O}_2$ the correct coefficient for oxygen would be:

 - i. 3

3. Types of reactions

- a. What are the different types of reactions?
 - i. Synthesis
 - ii. Decomposition
 - iii. Single replacement
 - iv. Double replacement
 - v. Combustion

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