

Oxidation and Reduction

What happens when electrons are transferred in a chemical reaction?

Why?

Silver tarnishes when it comes in contact with sulfur compounds in the air. Copper gets coated in beautiful green patina as it ages. Metals rust or corrode in the presence of air and water. Minerals (ionic compounds) found in ore can be decomposed with the use of electricity to produce pure metals and nonmetals. All of these reactions are examples of oxidation and reduction, otherwise known as **redox** reactions. In this activity you will explore what is happening at the atomic level in redox reactions.

Model 1 – Redox Reactions

Redox Reactions



Nonredox Reactions



1. What two types of reactions are shown in Model 1?
2. Examine the redox and nonredox reactions in Model 1. Is/are there any feature(s) in the redox reactions that would allow you to identify them as redox reactions? If yes, use specific examples from Model 1 to support your answer.

Model 1 Redox Reactions Answers Pogil

LL Leslie



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Chemistry Education in the ICT Age Minu Gupta Bhowon, Sabina Jhaumeer-Laulloo, Henri Li Kam Wah, Ponnadurai Ramasami, 2009-07-21 The 20 International Conference on Chemical Education 20 ICCE which had rd th Chemistry in the ICT Age as the theme was held from 3 to 8 August 2008 at Le Meridien Hotel Pointe aux Piments in Mauritius With more than 200 participants from 40 countries the conference featured 140 oral and 50 poster presentations th Participants of the 20 ICCE were invited to submit full papers and the latter were subjected to peer review The selected accepted papers are collected in this book of proceedings This book of proceedings encloses 39 presentations covering topics ranging from fundamental to applied chemistry such as Arts and Chemistry Education Biochemistry and Biotechnology Chemical Education for Development Chemistry at Secondary Level Chemistry at Tertiary Level Chemistry Teacher Education Chemistry and Society Chemistry Olympiad Context Oriented Chemistry ICT and Chemistry Education Green Chemistry Micro Scale Chemistry Modern Technologies in Chemistry Education Network for Chemistry and Chemical Engineering Education Public Understanding of Chemistry Research in Chemistry Education and Science Education at Elementary Level We would like to thank those who submitted the full papers and the reviewers for their timely help in assessing the papers for publication th We would also like to pay a special tribute to all the sponsors of the 20 ICCE and in particular the Tertiary Education Commission <http://tec.intnet.mu> and the Organisation for the Prohibition of Chemical Weapons <http://www.opcw.org> for kindly agreeing to fund the publication of these proceedings Redox Models in Chemistry Lise-Lotte Österlund, Margareta Ekborg, Anders Berg, 2010 *Redox Reactions Explained (General Chemistry Quick Review)* E Staff, Learn and review on the go Use Quick Review Science Study Notes to help you learn or brush up on the subject quickly You can use the review notes as a reference to understand the subject better and improve your grades Easy to remember facts to help you perform better Perfect study notes for all high school and college students **A Better Way to Learn Redox Chemistry Part 1: Summarized Principles and Over 250 Answered Questions and Problems** Malika Ammam, 2017-11-25 This short book part 1 summarizes some basics of redox chemistry electrochemistry in solution Each section of the book contains concise and clear referenced summaries of redox concepts Over 250 answered questions and problems are provided to further clarify the discussed principles *Investigation Into the Role of Redox Reactions in Maillard Model Systems* Luke John William Haffenden, 2007 The role of redox reactions in the formation of aroma volatiles colour and other non volatiles in the Maillard Reaction was investigated The electrochemical properties of individual reactants and Maillard model mixtures were monitored via ORP oxidation reduction potential and oxygen electrodes All models exhibited unique electrochemical activities represented by their corresponding ORP profiles Investigation into the redox potentials of several model systems demonstrated that the increased negative value of a redox potential is not necessarily correlated with its browning potential An optimal redox potential range where browning is favoured was found to represent a balance between carbonyl and

hydroxyl moieties in the structure Adjustment of this redox potential by introducing reducing or oxidizing species can shift this balance resulting in modifications in browning capacities However it was concluded that there is a clear relationship between browning ability and reducing capacity of the model systems Furthermore a novel oxidative pyrolysis technique was developed to study the role of oxidative environment on the product distribution during pyrolysis and to investigate the mechanism of formation of non volatiles through ^{13}C and ^{15}N label incorporation Application of this technique to glucoselglycine model system have indicated that most non volatile Maillard reaction products can arise from glucose oxidation intermediates such as glucosone gluconic acid and deoxyglucosones To study the specific role of redox reactions in the formation of non volatiles a post pyrolytic derivatization technique was developed and optimized Several non volatile end products were identified and mechanistically confirmed to involve oxidation and reduction reactions for their formation such as lactones hydroxylated benzenes and hydroxylated pyrazine The latter was identified and confirmed to be generated via the dimerization of glycine and subsequent oxidation In addition the formation of different volatiles such as pyrazole imidazole and oxazole was mechanistically confirmed to depend on redox reactions

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