

MATHEMATICS

Date: 21/ 06/2022

Period: 8:30-11:30



END OF TERM III EXAMINATIONS

GRADE
OPTION

SENIOR ONE
ORDINARY LEVEL

DURATION:

3HOURS

MARKS:

100

INSTRUCTIONS

1)This paper consists of **one** section

Section A: Attempt **all** questions.

(100marks)

2)You may use mathematical instruments and a calculator **where necessary**.

3)Use a **blue or black ink pen only** to write your answers and a **pencil** to draw diagrams.

4)Show clearly all the working steps. **Marks will not be awarded for the answer without all working steps.**

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Table of Contents Maths S1 18th January 2013 Mark Scheme

1. Understanding the eBook Maths S1 18th January 2013 Mark Scheme
 - The Rise of Digital Reading Maths S1 18th January 2013 Mark Scheme
 - Advantages of eBooks Over Traditional Books
2. Identifying Maths S1 18th January 2013 Mark Scheme
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Maths S1 18th January 2013 Mark Scheme
 - User-Friendly Interface
4. Exploring eBook Recommendations from Maths S1 18th January 2013 Mark Scheme
 - Personalized Recommendations
 - Maths S1 18th January 2013 Mark Scheme User Reviews and Ratings
 - Maths S1 18th January 2013 Mark Scheme and Bestseller Lists
5. Accessing Maths S1 18th January 2013 Mark Scheme Free and Paid eBooks

- Maths S1 18th January 2013 Mark Scheme Public Domain eBooks
 - Maths S1 18th January 2013 Mark Scheme eBook Subscription Services
 - Maths S1 18th January 2013 Mark Scheme Budget-Friendly Options
6. Navigating Maths S1 18th January 2013 Mark Scheme eBook Formats
- ePub, PDF, MOBI, and More
 - Maths S1 18th January 2013 Mark Scheme Compatibility with Devices
 - Maths S1 18th January 2013 Mark Scheme Enhanced eBook Features
7. Enhancing Your Reading Experience
- Adjustable Fonts and Text Sizes of Maths S1 18th January 2013 Mark Scheme
 - Highlighting and Note-Taking Maths S1 18th January 2013 Mark Scheme
 - Interactive Elements Maths S1 18th January 2013 Mark Scheme
8. Staying Engaged with Maths S1 18th January 2013 Mark Scheme
- Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Maths S1 18th January 2013 Mark Scheme
9. Balancing eBooks and Physical Books Maths S1 18th January 2013 Mark Scheme
- Benefits of a Digital Library
 - Creating a Diverse Reading Collection Maths S1 18th January 2013 Mark Scheme
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Maths S1 18th January 2013 Mark Scheme
- Setting Reading Goals Maths S1 18th January 2013 Mark Scheme
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Maths S1 18th January 2013 Mark Scheme
- Fact-Checking eBook Content of Maths S1 18th January 2013 Mark Scheme
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development

- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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solution the higher the ksp the more the chemical will dissolve into that solvent for this lab calcium hydroxide is created then
ionized Ca(OH)_2 when calcium is put in water reacts as follows $\text{Ca} + 2\text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2 + \text{H}_2$

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web k sp of calcium hydroxide 1 author j m mccormick last update ocober 13 2013 introduction an equilibrium constant k is
related to ΔG for a process through eqn 1 and as long as ΔH and ΔS are independent of temperature eqn 1 can be combined
with eqn 2 to determine ΔH and ΔS from the temperature dependence of k

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web in today's experiment you will determine the solubility product ksp of calcium hydroxide Ca(OH)_2 by measuring the concentration of Ca(OH)_2 in a saturated solution calcium hydroxide is a sparingly soluble salt that dissolves according to the following reaction the solubility product expression for this reaction is

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web expert verified chem 182 determining the ksp of calcium hydroxide calcium hydroxide is an ionic solid that is sparingly soluble in water a saturated aqueous solution of Ca(OH)_2 is represented in equation form as shown below $\text{Ca(OH)}_2(\text{s}) \rightleftharpoons \text{Ca}^{2+}(\text{aq}) + 2\text{OH}^{-}(\text{aq})$ the solubility product expression describes in mathematical terms the equilibrium

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web ksp for calcium hydroxide calcium hydroxide is a soft white caustic powder used in making mortar cements calcium salts paints and petrochemicals it is also used in saltwater aquaria to make up kalkwasser limewater solutions for reef tanks and is used as a pH regulating agent notice that calcium hydroxide is divalent and thus has twice

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web use the concentration of calcium hydroxide to calculate the concentration of calcium ions and hydroxide ions in the solution $\text{Ca(OH)}_2 \rightarrow \text{Ca}^{2+} + 2\text{OH}^{-}$ $5 \times 10^{-2} \text{ M}$ calcium ions $5 \times 10^{-2} \text{ M}$ $2 \times 10^{-2} \text{ M}$ hydroxide ions calculate calcium hydroxide ksp using the concentrations of calcium ions and hydroxide ions

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web jul 12 2023 because the concentration of a pure solid such as $\text{Ca}_3(\text{PO}_4)_2$ is a constant it does not appear explicitly in the equilibrium constant expression the equilibrium constant expression for the dissolution of calcium phosphate is therefore $K_{\text{sp}} \text{ Ca}_3(\text{PO}_4)_2$

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