

## Limiting Reactant and Percent Yield Worksheet

Key

1. Calcium hydroxide, used to neutralize acid spills, reacts with hydrochloric acid according to the following unbalanced equation:



- a. If you have spilled 6.3 mol of HCl and put 2.8 mol of  $\text{Ca(OH)}_2$  on it, which substance is the limiting reactant?

$$\frac{6.3 \text{ mol HCl}}{2 \text{ mol HCl}} \bigg| \frac{1 \text{ mol CaCl}_2}{2 \text{ mol HCl}} = 3.2 \text{ mol CaCl}_2$$

$$\frac{2.8 \text{ mol Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \bigg| \frac{1 \text{ mol CaCl}_2}{1 \text{ mol Ca(OH)}_2} = 2.8 \text{ mol CaCl}_2$$

\*  $\text{Ca(OH)}_2$  is limiting

- b. How many moles of the excess reactant remain?

$$\frac{2.8 \text{ mol Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \bigg| \frac{2 \text{ mol HCl}}{1 \text{ mol Ca(OH)}_2} = 5.6 \text{ mol HCl used}$$

$$6.3 \text{ mol HCl} - 5.6 \text{ mol HCl} = \boxed{0.7 \text{ mol HCl remain}}$$

2. Aluminum oxidizes according to the following unbalanced equation:  $\underline{4}\text{Al} + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{Al}_2\text{O}_3$

- a. Powdered Al (0.048 mol) is placed into a container containing 0.030 mol  $\text{O}_2$ . What is limiting reactant?

$$\frac{0.048 \text{ mol Al}}{4 \text{ mol Al}} \bigg| \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} = 0.024 \text{ mol Al}_2\text{O}_3$$

$$\frac{0.030 \text{ mol O}_2}{3 \text{ mol O}_2} \bigg| \frac{2 \text{ mol Al}_2\text{O}_3}{3 \text{ mol O}_2} = 0.020 \text{ mol Al}_2\text{O}_3$$

\*  $\text{O}_2$  is limiting

- b. How many moles of the excess reactant remain?

$$\frac{0.030 \text{ mol O}_2}{3 \text{ mol O}_2} \bigg| \frac{4 \text{ mol Al}}{3 \text{ mol O}_2} = 0.040 \text{ mol Al used}$$

$$0.048 \text{ mol Al} - 0.040 \text{ mol Al} = \boxed{0.008 \text{ mol Al remain}}$$

3. Heating zinc sulfide in the presence of oxygen yields the following:  $\underline{2}\text{ZnS} + \underline{3}\text{O}_2 \rightarrow \underline{2}\text{ZnO} + \underline{2}\text{SO}_2$

- a. If 1.72 mol of ZnS is heated in the presence of <sup>45.6</sup>~~32.0~~ g of  $\text{O}_2$ , which is the limiting reactant? (Balance the equation first.)

$$\frac{1.72 \text{ mol ZnS}}{1 \text{ mol ZnS}} \bigg| \frac{2 \text{ mol SO}_2}{1 \text{ mol ZnS}} = 3.44 \text{ mol SO}_2$$

$$\frac{45.6 \text{ g O}_2}{32.0 \text{ g O}_2} \bigg| \frac{1 \text{ mol O}_2}{32.0 \text{ g O}_2} \bigg| \frac{2 \text{ mol SO}_2}{3 \text{ mol O}_2} = 0.950 \text{ mol SO}_2$$

\*  $\text{O}_2$  is limiting

- b. How many grams of the excess reactant remain?

$$\frac{45.6 \text{ g O}_2}{32.0 \text{ g O}_2} \bigg| \frac{1 \text{ mol O}_2}{32.0 \text{ g O}_2} \bigg| \frac{2 \text{ mol ZnS}}{3 \text{ mol O}_2} \bigg| \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 92.6 \text{ g used}$$

$$\frac{1.72 \text{ mol ZnS}}{1 \text{ mol ZnS}} \bigg| \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 168 \text{ g ZnS (started with)}$$

$$\frac{168 \text{ g ZnS} - 92.6 \text{ g ZnS}}{1} = \boxed{75 \text{ g ZnS remain}}$$

# Limiting Reagent And Percent Yield Answer Key 12 3

**Katrin Zwirgmaier**



## **Limiting Reagent And Percent Yield Answer Key 12 3:**

## **Limiting Reagent And Percent Yield Answer Key 12 3** Book Review: Unveiling the Power of Words

In a world driven by information and connectivity, the ability of words has become more evident than ever. They have the capability to inspire, provoke, and ignite change. Such could be the essence of the book **Limiting Reagent And Percent Yield Answer Key 12 3**, a literary masterpiece that delves deep into the significance of words and their affect our lives. Compiled by a renowned author, this captivating work takes readers on a transformative journey, unraveling the secrets and potential behind every word. In this review, we shall explore the book's key themes, examine its writing style, and analyze its overall affect readers.

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