

## Limiting Reactant and Percent Yield Worksheet

Key

1. Calcium hydroxide, used to neutralize acid spills, reacts with hydrochloric acid according to the following unbalanced equation:



- a. If you have spilled 6.3 mol of HCl and put 2.8 mol of  $\text{Ca(OH)}_2$  on it, which substance is the limiting reactant?

$$\frac{6.3 \text{ mol HCl}}{2 \text{ mol HCl}} \bigg| \frac{1 \text{ mol CaCl}_2}{2 \text{ mol HCl}} = 3.2 \text{ mol CaCl}_2$$

$$\frac{2.8 \text{ mol Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \bigg| \frac{1 \text{ mol CaCl}_2}{1 \text{ mol Ca(OH)}_2} = 2.8 \text{ mol CaCl}_2$$

\*  $\text{Ca(OH)}_2$  is limiting

- b. How many moles of the excess reactant remain?

$$\frac{2.8 \text{ mol Ca(OH)}_2}{1 \text{ mol Ca(OH)}_2} \bigg| \frac{2 \text{ mol HCl}}{1 \text{ mol Ca(OH)}_2} = 5.6 \text{ mol HCl used}$$

$$6.3 \text{ mol HCl} - 5.6 \text{ mol HCl} = \boxed{0.7 \text{ mol HCl remain}}$$

2. Aluminum oxidizes according to the following unbalanced equation:  $4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$

- a. Powdered Al (0.048 mol) is placed into a container containing 0.030 mol  $\text{O}_2$ . What is limiting reactant?

$$\frac{0.048 \text{ mol Al}}{4 \text{ mol Al}} \bigg| \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} = 0.024 \text{ mol Al}_2\text{O}_3$$

$$\frac{0.030 \text{ mol O}_2}{3 \text{ mol O}_2} \bigg| \frac{2 \text{ mol Al}_2\text{O}_3}{3 \text{ mol O}_2} = 0.020 \text{ mol Al}_2\text{O}_3$$

\*  $\text{O}_2$  is limiting

- b. How many moles of the excess reactant remain?

$$\frac{0.030 \text{ mol O}_2}{3 \text{ mol O}_2} \bigg| \frac{4 \text{ mol Al}}{3 \text{ mol O}_2} = 0.040 \text{ mol Al used}$$

$$0.048 \text{ mol Al} - 0.040 \text{ mol Al} = \boxed{0.008 \text{ mol Al remain}}$$

3. Heating zinc sulfide in the presence of oxygen yields the following:  $2\text{ZnS} + 3\text{O}_2 \rightarrow 2\text{ZnO} + 2\text{SO}_2$

- a. If 1.72 mol of ZnS is heated in the presence of <sup>45.6</sup>~~32.0~~ g of  $\text{O}_2$ , which is the limiting reactant? (Balance the equation first.)

$$\frac{1.72 \text{ mol ZnS}}{1 \text{ mol ZnS}} \bigg| \frac{2 \text{ mol SO}_2}{1 \text{ mol ZnS}} = 3.44 \text{ mol SO}_2$$

$$\frac{45.6 \text{ g O}_2}{32.0 \text{ g O}_2} \bigg| \frac{1 \text{ mol O}_2}{32.0 \text{ g O}_2} \bigg| \frac{2 \text{ mol SO}_2}{3 \text{ mol O}_2} = 0.450 \text{ mol SO}_2$$

\*  $\text{O}_2$  is limiting

- b. How many grams of the excess reactant remain?

$$\frac{45.6 \text{ g O}_2}{32.0 \text{ g O}_2} \bigg| \frac{1 \text{ mol O}_2}{32.0 \text{ g O}_2} \bigg| \frac{2 \text{ mol ZnS}}{3 \text{ mol O}_2} \bigg| \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 92.6 \text{ g used}$$

$$\frac{1.72 \text{ mol ZnS}}{1 \text{ mol ZnS}} \bigg| \frac{97.48 \text{ g ZnS}}{1 \text{ mol ZnS}} = 168 \text{ g ZnS (started with)}$$

$$\begin{array}{r} 168 \text{ g ZnS} \\ - 92.6 \text{ g ZnS} \\ \hline 75 \text{ g ZnS} \\ \text{remain} \end{array}$$

# Limiting Reagent And Percent Yield 12 3 Answers

**Sara McCubbins, Angela Codron**



## **Limiting Reagent And Percent Yield 12 3 Answers:**

CHEMISTRY HANDBOOK & STUDY GUIDE Gr11-12 NE Kevin Smith, 2024-02-01 A comprehensive summary of Grade 11 12 Physics Simple logical summaries with example exam questions and work through solutions The book covers the fundamentals of Grade 11 12 Physics and complements the material in any class text Class 8-12 Chemistry Questions and Answers PDF Arshad Iqbal, The Class 8 12 Chemistry Quiz Questions and Answers PDF Grade 8 12 Chemistry Competitive Exam Questions Chapter 1 15 Practice Tests Chemistry Textbook Questions for Beginners includes Questions to solve problems with hundreds of class questions Class 8 12 Chemistry Questions and Answers PDF book covers basic concepts and analytical assessment tests Class 8 12 Chemistry Quiz PDF book helps to practice test questions from exam prep notes The Grade 8 12 Chemistry Quiz Questions and Answers PDF eBook includes Practice material with verbal quantitative and analytical past papers questions Class 8 12 Chemistry Questions and Answers PDF Free download chapter 1 a book to review textbook questions on chapters Molecular structure acids and bases atomic structure bonding chemical equations descriptive chemistry equilibrium systems gases laboratory liquids and solids mole concept oxidation reduction rates of reactions solutions thermochemistry Questions for high school and college revision questions Chemistry Interview Questions and Answers PDF Download free eBook s sample covers beginner s solved questions textbook s study notes to practice online tests The Grade 8 12 Chemistry Interview Questions Chapter 1 15 PDF book includes high school workbook questions to practice Questions for exam Chemistry Practice Tests a textbook s revision guide with chapters Questions for NEET MCAT GRE GMAT SAT ACT competitive exam Grade 8 12 Chemistry Questions Bank Chapter 1 15 PDF book covers problem solving exam tests from chemistry practical and textbook s chapters as Chapter 1 Molecular Structure Questions Chapter 2 Acids and Bases Questions Chapter 3 Atomic Structure Questions Chapter 4 Bonding Questions Chapter 5 Chemical Equations Questions Chapter 6 Descriptive Chemistry Questions Chapter 7 Equilibrium Systems Questions Chapter 8 Gases Questions Chapter 9 Laboratory Questions Chapter 10 Liquids and Solids Questions Chapter 11 Mole Concept Questions Chapter 12 Oxidation Reduction Questions Chapter 13 Rates of Reactions Questions Chapter 14 Solutions Questions Chapter 15 Thermochemistry Questions The Molecular Structure Quiz Questions PDF e Book Chapter 1 interview questions and answers on polarity three dimensional molecular shapes The Acids and Bases Quiz Questions PDF e Book Chapter 2 interview questions and answers on Arrhenius concept Bronsted lowry concept indicators introduction Lewis concept pH strong and weak acids and bases The Atomic Structure Quiz Questions PDF e Book Chapter 3 interview questions and answers on electron configurations experimental evidence of atomic structure periodic trends quantum numbers and energy levels The Bonding Quiz Questions PDF e Book Chapter 4 interview questions and answers on ionic bond covalent bond dipole dipole forces hydrogen bonding intermolecular forces London dispersion forces metallic bond The Chemical Equations Quiz Questions PDF e Book Chapter 5 interview questions and answers on balancing of equations limiting reactants percent yield

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Chemistry students for the required laboratory experiments section on the exam      **Quick Access** Research and Education Association, 2009-09-25 REA's Quick Access Study Charts contain all the information students, teachers, and professionals need in one handy reference. They provide quick, easy access to important facts. The charts contain commonly used math formulas, historical facts, language conjugations, vocabulary, and more. Great for exams, classroom reference, or a quick refresher on the subject.

**Chemical Principles** Peter Atkins, Loretta Jones, 2007-08 Written for calculus-inclusive general chemistry courses, *Chemical Principles* helps students develop chemical insight by showing the connections between fundamental chemical ideas and their applications. Unlike other texts, it begins with a detailed picture of the atom, then builds toward chemistry's frontier, continually demonstrating how to solve problems, think about nature and matter, and visualize chemical concepts as working chemists do. Flexibility in level is crucial and is largely established through clearly labeling separating in boxes the calculus coverage in the text. Instructors have the option of whether to incorporate calculus in the coverage of topics. The multimedia integration of *Chemical Principles* is more deeply established than any other text for this course. Through the unique eBook, the comprehensive Chemistry Portal, Living Graph icons that connect the text to the Web, and a complete set of animations, students can take full advantage of the wealth of resources available to them to help them learn and gain a deeper understanding.

Concepts And Problems In Physical Chemistry P.S. Raghavan, 1997 Contents: Introduction, Atoms, Molecules, and Formulas, Chemical Equations and Stoichiometry, Aqueous Reactions and Solution Stoichiometry, Gases, Intermolecular Forces, Liquids and Solids, Atoms, Structure and the Periodic Table, Chemical Bonding, Chemical Thermodynamics, Solutions, Chemical Kinetics, Chemical Equilibrium, Acids and Bases, Ionic Equilibria I, Ionic Equilibria II, Redox Reactions, Electrochemistry, Nuclear Chemistry.

**Physical Science** C. Lon Enloe, Elizabeth Garnett, Jonathan Miles, Stephen Swanson, 2000-12-25 This is an introductory book that provides students with the tools to master the basic principles of physics and chemistry needed by the aspiring technology professional. Like all the books in the critically acclaimed *Preserving the Legacy* series, each chapter is divided into subsections featuring learning objectives and a *Check Your Understanding* section to help students focus on important concepts. Questions requiring written and mathematical answers at the end of each chapter provide students with the opportunity to further demonstrate their understanding of the concepts. The only book available that specifically addresses the emerging need for a course to teach physics and chemistry principles to the growing number of students entering the various fields of technology, it offers a thorough grounding in foundational concepts along with Technology boxes that offer practical applications.

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