

# Sample Problem 1 Worked Out



Step 1 Find the limiting reagent:

$$5.22 \text{ kg H}_2 \times \frac{1000 \text{ g H}_2}{1 \text{ kg H}_2} \times \frac{1 \text{ mol H}_2}{2.02 \text{ g H}_2} \times \frac{2 \text{ mol NH}_3}{3 \text{ mol H}_2} = 1.72 \times 10^3 \text{ mol NH}_3$$

$$31.5 \text{ kg N}_2 \times \frac{1000 \text{ g N}_2}{1 \text{ kg N}_2} \times \frac{1 \text{ mol N}_2}{28.02 \text{ g N}_2} \times \frac{2 \text{ mol NH}_3}{1 \text{ mol N}_2} = 2.25 \times 10^3 \text{ mol NH}_3$$

H<sub>2</sub> is the limiting reagent

# Limiting Reagent And Percent Yield 12 3

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