

## Limiting Reactant Worksheet #2

1. Given the following reaction: (Balance the equation first!)



- If you start with 14.8 g of  $\text{C}_3\text{H}_8$  and 3.44 g of  $\text{O}_2$ , determine the limiting reactant.
- Determine the number of moles of carbon dioxide produced.
- Determine the number of grams of  $\text{H}_2\text{O}$  produced.
- Determine the number of grams of excess reactant left.

2. Given the following equation:



- If 10.0 g of  $\text{Al}_2(\text{SO}_4)_3$  is reacted with 10.0 g of  $\text{NaOH}$ , determine the limiting reactant.
- Determine the number of moles of  $\text{Al}(\text{OH})_3$  produced.
- Determine the number of grams of  $\text{Na}_2\text{SO}_4$  produced.
- Determine the number of grams of excess reactant left over in the reaction.

3. Given the following equation:



- If 25.4 g of  $\text{Al}_2\text{O}_3$  is reacted with 10.2 g of  $\text{Fe}$ , determine the limiting reactant.
- Determine the number of moles of  $\text{Al}$  produced.
- Determine the number of grams of  $\text{Fe}_3\text{O}_4$  produced.
- Determine the number of grams of excess reactant left over in the reaction.

4. Consider the reaction



- 80.0 grams of  $\text{I}_2\text{O}_5$ , reacts with 28.0 grams of carbon monoxide,  $\text{CO}$ . Determine the mass of iodine  $\text{I}_2$ , which could be produced?
- If, in the above situation, only 0.160 moles, of iodine,  $\text{I}_2$  was produced.
  - What mass of iodine was produced?
  - What percentage yield of iodine was produced?

5. Zinc and sulfur react to form zinc sulfide according to the equation



If 25.0 g of zinc and 30.0 g of sulfur are mixed,

- Which chemical is the limiting reactant?
- How many grams of  $\text{ZnS}$  will be formed?
- How many grams of the excess reactant will remain after the reaction is over?

6. Which element is in excess when 3.00 grams of  $\text{Mg}$  is ignited in 2.20 grams of pure oxygen? What mass is in excess? What mass of  $\text{MgO}$  is formed?

7. How many grams of  $\text{Al}_2\text{S}_3$  are formed when 5.00 grams of  $\text{Al}$  is heated with 10.0 grams  $\text{S}$ ?

# Limiting Reagent Practice 2 Answers

**RD Boyd**



## Limiting Reagent Practice 2 Answers:

The Practice of Chemistry Study Guide & Solutions Manual Pamela Mills, Amina El-Ashmawy, 2003-04-14 Designed to help students understand the material better and avoid common mistakes Also includes solutions and explanations to odd numbered exercises     **The Official ACT Prep Guide 2019-2020, (Book + 5 Practice Tests + Bonus Online Content)**

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Day 10 Ionic Equilibrium

Day 11 Unit Test 2 Physical Chemistry I

Day 12 Redox Reactions

Day 13 Electrochemistry

Day 14 Chemical Kinetics

Day 15 Adsorption and Catalysis

Day 16 Colloidal State

Day 17 Unit Test 3 Physical Chemistry II

Day 18 Classification and Periodicity of Elements

Day 19 General Principles and Processes of Isolation of Metals

Day 20 Hydrogen

Day 21 s Block Elements

Day 22 p Block Elements Group 13 to Group 18

Day 23 The d and f Block Elements

Day 24 Coordination Compounds

Day 25 Unit Test 4 Inorganic Chemistry

Day 26 Environmental Chemistry

Day 27 General Organic Chemistry

Day 28 Hydrocarbons

Day 29 Organic Compounds Containing Halogens

Day 30 Organic Compounds Containing Oxygen

Day 31 Organic Compounds Containing Nitrogen

Day 32 Unit Test 5 Organic Chemistry I

Day 33 Polymers

Day 34 Biomolecules

Day 35 Chemistry in Everyday Life

Day 36 Analytical Chemistry

Day 37 Unit Test 6 Organic Chemistry II

Day 38 Mock Test 1

Day 39 Mock Test 2

Day 40 Mock Test 3

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## **Table of Contents Limiting Reagent Practice 2 Answers**

1. Understanding the eBook Limiting Reagent Practice 2 Answers
  - The Rise of Digital Reading Limiting Reagent Practice 2 Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reagent Practice 2 Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Limiting Reagent Practice 2 Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reagent Practice 2 Answers
  - Personalized Recommendations
  - Limiting Reagent Practice 2 Answers User Reviews and Ratings
  - Limiting Reagent Practice 2 Answers and Bestseller Lists
5. Accessing Limiting Reagent Practice 2 Answers Free and Paid eBooks
  - Limiting Reagent Practice 2 Answers Public Domain eBooks
  - Limiting Reagent Practice 2 Answers eBook Subscription Services
  - Limiting Reagent Practice 2 Answers Budget-Friendly Options
6. Navigating Limiting Reagent Practice 2 Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Limiting Reagent Practice 2 Answers Compatibility with Devices
  - Limiting Reagent Practice 2 Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Limiting Reagent Practice 2 Answers
  - Highlighting and Note-Taking Limiting Reagent Practice 2 Answers
  - Interactive Elements Limiting Reagent Practice 2 Answers
8. Staying Engaged with Limiting Reagent Practice 2 Answers

- Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers
9. Balancing eBooks and Physical Books
- Benefits of a Digital Library
  - Creating a Diverse Reading Collection
10. Overcoming Reading Challenges
- Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine
- Setting Reading Goals
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information
- Fact-Checking eBook Content
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning
- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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