Limiting Reactant Worksheet #2

1. Given the following reaction: (Balance the equation first!)

C3H8 + O2 -----> CO2 + H2O

- a) If you start with 14.8 g of C₃H₈ and 3.44 g of O₂, determine the limiting reactant.
- b) Determine the number of moles of carbon dioxide produced.
- c) Determine the number of grams of H2O produced.
- d) Determine the number of grams of excess reactant left.
- Given the following equation:

Al₂(SO₃)₃ + 6NaOH -----> 3Na₂SO₃ + 2Al(OH)₃

- a) If 10.0 g of Al₂(SO₃)₃ is reacted with 10.0 g of NaOH, determine the limiting reactant.
- b) Determine the number of moles of Al(OH)₃ produced.
- c) Determine the number of grams of Na₂SO₃ produced.
- d) Determine the number of grams of excess reactant left over in the reaction.
- 3. Given the following equation:

 Al_2O_3 + Fe ----> Fe_3O_4 + Al

- a) If 25.4 g of Al₂O₃ is reacted with 10.2 g of Fe, determine the limiting reactant.
- b) Determine the number of moles of Al produced.
- c) Determine the number of grams of Fe₃O₄ produced.
- d) Determine the number of grams of excess reactant left over in the reaction.
- Consider the reaction

 $I_2O_5(g) + 5 CO(g) -----> 5 CO_2(g) + I_2(g)$

- a) 80.0 grams of I2O5, reacts with 28.0 grams of carbon monoxide, CO.
- Determine the mass of iodine I₂, which could be produced?

 b) If, in the above situation, only 0.160 moles, of iodine, I₂ was produced.
 - i) What mass of iodine was produced?
 - ii) What percentage yield of iodine was produced?
- 5. Zinc and sulfur react to form zinc sulfide according to the equation

Zn + S -----> ZnS

If 25.0 g of zinc and 30.0 g of sulfur are mixed,

- a) Which chemical is the limiting reactant?
- b) How many grams of ZnS will be formed?
- c) How many grams of the excess reactant will remain after the reaction is over?
- 6. Which element is in excess when 3.00 grams of Mg is ignited in 2.20 grams of pure oxygen? What mass is in excess? What mass of MgO is formed?
- 7. How many grams of Al₂S₃ are formed when 5.00 grams of Al is heated with 10.0 grams S?

Limiting Reagent Answers And Work

Neil D. Jespersen, Pamela Kerrigan

Limiting Reagent Answers And Work:

Unveiling the Power of Verbal Art: An Emotional Sojourn through Limiting Reagent Answers And Work

In a world inundated with screens and the cacophony of quick communication, the profound power and mental resonance of verbal art usually fade in to obscurity, eclipsed by the constant onslaught of noise and distractions. However, situated within the musical pages of **Limiting Reagent Answers And Work**, a fascinating function of literary beauty that impulses with raw thoughts, lies an memorable trip waiting to be embarked upon. Composed with a virtuoso wordsmith, this interesting opus manuals visitors on an emotional odyssey, lightly revealing the latent possible and profound influence embedded within the delicate internet of language. Within the heart-wrenching expanse of the evocative analysis, we can embark upon an introspective exploration of the book is key themes, dissect their interesting publishing model, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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