

Limiting Reactant Practice 2.3 - Answer Key

Objective: Use stoichiometry to identify the limiting and excess reactants.

Directions: Answer the following limiting reactant problems. Show your work!!

1. A reaction between potassium and oxygen gas produces potassium oxide, represented by the following balanced chemical equation:



- a. If 3.00 g of potassium and 5.00 g of oxygen are mixed, how much potassium oxide is produced?

$$3.00 \text{ g K} \times \left(\frac{1 \text{ mol K}}{39.10 \text{ g K}} \right) \left(\frac{1 \text{ mol O}_2}{2 \text{ mol K}} \right) \left(\frac{2 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) = \boxed{1.21 \text{ g K}_2\text{O}}$$

$$5.00 \text{ g O}_2 \times \left(\frac{1 \text{ mol O}_2}{32.00 \text{ g O}_2} \right) \left(\frac{1 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) \left(\frac{2 \text{ mol K}_2\text{O}}{1 \text{ mol O}_2} \right) = 3.12 \text{ g K}_2\text{O}$$

- b. What is the limiting reactant?

Potassium (K)

- c. What is the reactant in excess?

Oxygen (O₂)

2. In a reaction, 12.00 g of aqueous sodium hydroxide reacts with aqueous 22.00 g of copper(II) sulfate.



- a. What mass of sodium sulfate is produced?

$$12.00 \text{ g NaOH} \times \left(\frac{1 \text{ mol NaOH}}{40.00 \text{ g NaOH}} \right) \left(\frac{1 \text{ mol Na}_2\text{SO}_4}{2 \text{ mol NaOH}} \right) \left(\frac{142.04 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} \right) = 17.10 \text{ g}$$

$$22.00 \text{ g CuSO}_4 \times \left(\frac{1 \text{ mol CuSO}_4}{159.61 \text{ g CuSO}_4} \right) \left(\frac{1 \text{ mol Na}_2\text{SO}_4}{1 \text{ mol CuSO}_4} \right) \left(\frac{142.04 \text{ g Na}_2\text{SO}_4}{1 \text{ mol Na}_2\text{SO}_4} \right) = 19.87 \text{ g}$$

- b. What is the limiting reactant?

Copper(II) sulfate (CuSO₄)

- c. What is the reactant in excess?

Sodium Hydroxide (NaOH)

Limiting Reactants Study Guide Answer Key

DK



Limiting Reactants Study Guide Answer Key:

Basic Concepts of Chemistry, 9e Study Guide and Solutions Manual Leo J. Malone, Theodore O. Dolter, 2012-01-03 The 9th edition of Malone's Basic Concepts of Chemistry provides many new and advanced features that continue to address general chemistry topics with an emphasis on outcomes assessment. New and advanced features include an objectives grid at the end of each chapter which ties the objectives to examples within the sections, assessment exercises at the end of each section, and relevant chapter problems at the end of each chapter. A new Math Check allows quick access to the needed basic skill. The first chapter now includes brief introductions to several fundamental chemical concepts, and Chapter Synthesis Problems have been added to the end of each chapter to bring key concepts into one encompassing problem. Every concept in the text is clearly illustrated with one or more step-by-step examples. Making it Real essays have been updated to present timely and engaging real-world applications, emphasizing the relevance of the material they are learning. This edition continues the end-of-chapter Student Workshop activities to cater to the many different learning styles and to engage users in the practical aspect of the material discussed in the chapter. *Chemistry, Student Study Guide* James E. Brady, Fred Senese, 2008-01-28

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