
LIMITING REACTANT LAB

MATERIALS: 250 mL beaker, 100 mL graduated cylinder, spoon or stirring rod, and a watch glass.

SAFETY PRECAUTIONS: Wear goggles during this experiment. Avoid skin contact with copper (II) chloride dihydrate. Wash your hands before leaving the laboratory.

PRELAB QUESTION:

Aluminum metal reacts with aqueous copper (II) chloride dihydrate ($\text{CuCl}_2 \cdot 2 \text{H}_2\text{O}$) to form copper metal, aqueous aluminum chloride, and water. Write the balanced equation for this reaction, including state symbols.

PROCEDURE:

DAY 1

1. Weigh out the copper (II) chloride dihydrate. Record the true mass of the hydrated compound used.
2. Using the chart below, select your group number and the approximate masses of aluminum metal and copper (II) chloride dihydrate you'll be using in this activity:

Groups	Mass of Al (g)	Mass of $\text{CuCl}_2 \cdot 2 \text{H}_2\text{O}$ (g)
1 & 7 & 13	1.00	2.00
2 & 8 & 14	0.50	2.00
3 & 9	0.50	4.00
4 & 10	0.50	6.00
5 & 11	0.25	5.00
6 & 12	0.25	4.00

3. Using a graduated cylinder, measure water (tap water is fine) and add it to your LABELED beaker containing the copper (II) chloride dihydrate. Stir until the salt dissolves completely. Keep adding water (maximum of 100 mL) until all salt is dissolved – record the amount of water used. Note the color of the solution.
4. Weigh out the appropriate mass of aluminum foil, record that mass, then crumple it up into a few small pieces (NOT TIGHTLY CRUMPLED) and add it to the copper (II) chloride dihydrate solution. Stir the solution gently and then allow it to sit undisturbed for 5-10 minutes. Write observations.

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VM Jensen



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