

\_\_\_\_\_

100

- **Long-term effects**
- **Small but positive percentage of the 17000 people who took part in the study**
- **Percentage of people who were severely injured, hospitalized, or killed, and who lost consciousness or died**
- **People who were severely injured**
- **People who were hospitalized**
- **People who were killed**
- **People who lost consciousness**
- **People who died**
- **People who were severely injured, hospitalized, or killed**
- **People who lost consciousness**

Year	Number of cases	Number of deaths
1990	100	10
1991	120	12
1992	150	15
1993	180	18
1994	200	20
1995	220	22
1996	250	25
1997	280	28
1998	300	30
1999	320	32
2000	350	35
2001	380	38
2002	400	40
2003	420	42
2004	450	45
2005	480	48
2006	500	50
2007	520	52
2008	550	55
2009	580	58
2010	600	60
2011	620	62
2012	650	65
2013	680	68
2014	700	70
2015	720	72
2016	750	75
2017	780	78
2018	800	80
2019	820	82
2020	850	85
2021	880	88
2022	900	90
2023	920	92
2024	950	95
2025	980	98
2026	1000	100
2027	1020	102
2028	1050	105
2029	1080	108
2030	1100	110

100

Item	Quantity
1. 1000 1/2" x 1/2" x 1/2" Steel Balls	1
2. 1000 1/2" x 1/2" x 1/2" Steel Balls	1
3. 1000 1/2" x 1/2" x 1/2" Steel Balls	1
4. 1000 1/2" x 1/2" x 1/2" Steel Balls	1
5. 1000 1/2" x 1/2" x 1/2" Steel Balls	1
6. 1000 1/2" x 1/2" x 1/2" Steel Balls	1
7. 1000 1/2" x 1/2" x 1/2" Steel Balls	1
8. 1000 1/2" x 1/2" x 1/2" Steel Balls	1
9. 1000 1/2" x 1/2" x 1/2" Steel Balls	1
10. 1000 1/2" x 1/2" x 1/2" Steel Balls	1

[illegible]

\_\_\_\_\_

100

1000

[illegible]

1. The first step is to identify the problem or question that needs to be answered. This involves understanding the context and the specific requirements of the task.

# Lg Gh20 User Guide

**Rachel Sandford**



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web the pressure of the gas the pressure of the gas was calculated using dalton s law of partial pressures by subtracting the pressure of the water vapor from the total pressure which in this case was the atmospheric pressure 760mm hg the pressure of the water vapor was found by using the temperature of the water

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web in this lab you will collect a measurable volume of butane gas using water displacement using dalton s law and the ideal gas law the number of moles of butane gas and then the molar mass of butane can be calculated your experimental value for molar mass will be compared to the known molar mass of butane and percent error calculated

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web use the combined gas law to determine the volume in l of butane at stp use avogadro s law to determine the number of moles of butane gas assume that butane is an ideal gas and that one mole has a volume of 22 l at stp experimental molar mass g mole of butane the molecular formula of butane is c 4 h 10

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web although the most convenient gas for this is butane other gases may be available in similar small cylinders the collection of a gas sample and the weighing of the gas cylinder before and after this should take about 5 10 minutes as a demonstration

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