

5.1**MODELS OF THE ATOM****Section Review****Objectives**

- Identify inadequacies in the Rutherford atomic model
- Identify the new assumption in the Bohr model of the atom
- Describe the energies and positions of electrons according to the quantum mechanical model
- Describe how the shapes of orbitals at different sublevels vary

Vocabulary

- energy levels
- quantum
- quantum mechanical model
- atomic orbital

Part A Completion

Use this completion exercise to check your understanding of the concepts and terms that are introduced in this section. Each blank can be completed with a term, short phrase, or number.

- The chemical properties of atoms, ions, and molecules are related to the arrangement of the 1 within them. 1. _____
- The first modern atomic theory, proposed by 2, portrayed the atom as a solid, indivisible mass. After the discovery of the electron by 3, the atomic model was revised to include them. J.J. Thomson's model is referred to as the 4 model. Rutherford pictured the atom as a dense 5 surrounded by electrons. In the Bohr model, the electrons move in 6 paths. The 7 model is the modern description of the electrons in atoms. This model estimates the 8 of finding an electron within a certain volume of space surrounding the nucleus. 2. _____
3. _____
4. _____
5. _____
6. _____
7. _____
8. _____

Part B True-False

Classify each of these statements as always true, AT; sometimes true, ST; or never true, NT.

- _____ 9. Electrons must have a certain minimum amount of energy called a quantum in order to move from one energy level to the next higher energy level.
- _____ 10. The electron probability clouds for atomic orbitals are spherical in shape.

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N Colangelo



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