# 8 • Electron Configurations & Periodicity

#### WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

1.	2.	3.	4.	5.	6.
	, <u>S</u>				S S S S S S S S S S S S S S S S S S S
2000 O	200000 C	2000	2000	2 × × × × × × × × × × × × × × × × × × ×	80808 80808
¹	¹⊠	180	180	'₩	'₩
Element: Ar # of e 's: 16	Element: Mg # of e-'s: 12	Element: N # of e-'s: 7	Element: Li # of e-'s: _3	Element: P # of e 's: 15	Element: Cl # of e <sup>-</sup> 's: 17

Write the electron configurations of each of these in long form and short form:

- Ar 152 252 206 352 306
  - AT TNET 352 3pt
- Mg 152 252 206 352
  - Mg [Ne] 352
- N 152 252 2p3 N [He] 252 2p3
- Li 152 25'
- P 152 252 Zpc 352 3p3
  - P [Ne] 35+ 3p3
- CI 152 252 2p6 352 3p5
- CI [Ne] 352 305



Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.

Fe 1s2 252 2p 352 3p6 3d6 452

Fe [Ar] 3d 452



# **Electron Confuguration And Periodicity Lab Answer Key**

**M** Walker

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