



Electronic Configuration



1. Use noble gases to write the abbreviated electron configurations of the following?

i. Li _____

ii. B _____

iii. O _____

iv. Mg _____

v. Cl _____

vi. V _____

vii. Se^{2-} _____

viii. Cr^{3+} _____

ix. I _____

x. Ag _____

2. Identify the element with the following electron configuration

i. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^4$ _____

ii. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2$ _____

iii. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^7$ _____

iv. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10}$ _____

v. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^1$ _____

3. Determine if the following electron configurations are correct:

i. $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^6 5s^1$ _____

ii. $1s^2 2s^2 2p^6 3s^3$ _____

iii. $[\text{Rn}] 7s^2 5f^9 6d^2$ _____

iv. $[\text{Ar}] 5s^2 4d^{10} 5p^5$ _____

v. $[\text{Xe}] 6s^2 4f^{10}$ _____

4. Identify the following elements

i. $1s^2 2s^2 2p^2$ _____

ii. $1s^2 2s^2 2p^6$ _____

iii. $[\text{Ar}] 4s^2 3d^{10} 4p^5$ _____

iv. $[\text{Kr}] 5s^2 4d^1$ _____

v. $[\text{Ne}] 3s^2 3p^3$ _____

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Julius Sumner Miller



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