

1. Consider the following electron configurations:

Q $1s^2 2s^2 3s^2$

R $1s^2 2s^2 2p^6$

S $s^2 2s^2 2p^6 3s^1$

T $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^8$

V $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6$

Which configurations represent elements in the same group on the periodic table?
 A) R and V) B) Q and T) C) S and V) D) R and S)

ANS: C

1. What is the electron configuration of the sulfide ion, S^{2-} ?

A) $1s^2 2s^2 2p^6 3s^2 3p^4$ B) $1s^2 2s^2 2p^6 3s^2 3p^6$
 C) $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2$ D) $1s^2 2s^2 2p^6 3s^2 3p^3$

ANS: B

1. $[Ar]4s^2 3d^7 4p^1$ is the electron configuration of which _____ atom.

A) As B) V C) P D) Sb E) Sn

ANS: A

1. There are _____ unpaired electrons in a ground state phosphorus atom.

A) 0 B) 1 C) 2 D) 3 E) 4

ANS: D

1. There are _____ unpaired electrons in a ground state fluorine atom.

A) 0 B) 1 C) 2 D) 3 E) 4

ANS: B

1. In a ground-state manganese atom, the _____ subshell is partially filled.

A) 3s B) 4s C) 4p D) 3d E) 4d

ANS: D

1. The electron configuration of a ground-state Ag atom is _____.

A) $[Ar]4s^2 4d^9$ B) $[Kr]5s^2 4d^9$ C) $[Kr]4s^2 4d^9$ D) $[Ar]4s^2 4d^9$ E) $[Kr]5s^2 4d^9$

ANS: B

1. The ground state electron configuration for Zn is _____.

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