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Extension Questions

- 24. The hydrogen spectral lines in Model 2 are only the wavelengths of light that are in the visible range and therefore "seen" by the naked eye. However, many other wavelengths can be detected with special equipment.
 - a. Propose a hydrogen electron transition that involves light with a wavelength in the ultraviolet (UV) range (10–400 nm).

wavelengths within the UV range would have more energy than violet wavelengths. This transition would be from n=6 to n=2 or may be n=6 to n=1.

 Propose a hydrogen electron transition that involves light with a wavelength in the infrared (IR) range (1000–106 nm).

wavelengths within the IR range would have less energy than red wavelengths. This transition would be from n=3 to n=2 or maybe u=5 to n=4.

 Below are diagrams for the bright line spectra of four elements and the spectrum of a mixture of unknown gases.

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DJ Losen

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