

## Electron Configurations - KEY

*Note: The electron configurations in this worksheet assume that lanthanum (La) is the first element in the 4f block and that actinium (Ac) is the first element in the 5f block. If your periodic table doesn't agree with this, your answers for elements near the f-orbitals may be slightly different (one more/less electron for the f-orbital).*

- 1) sodium  $1s^2 2s^2 2p^6 3s^1$
- 2) iron  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^6$
- 3) bromine  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^5$
- 4) barium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2$
- 5) neptunium  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 5f^5$   
 also might be written (due to the close energy of d and f and similarities of elements versus predictable memorized config:  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^{10} 5p^6 6s^2 4f^{14} 5d^{10} 6p^6 7s^2 6d^1 5f^4$
- 6) cobalt  $[Ar] 4s^2 3d^7$
- 7) silver  $[Kr] 5s^2 4d^9$
- 8) tellurium  $[Kr] 5s^2 4d^{10} 5p^4$
- 9) radium  $[Rn] 7s^2$
- 10) lawrencium  $[Rn] 7s^2 5f^{14} 6d^1$   
 also might be written (due to the close energy of d and f and similarities of elements versus predictable memorized config:  $[Rn] 7s^2 6d^1 5f^{14}$
- 11)  $1s^2 2s^2 2p^6 3s^2 3p^4$  sulfur
- 12)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$  rubidium
- 13)  $[Kr] 5s^2 4d^{10} 5p^3$  antimony
- 14)  $[Xe] 6s^2 4f^{14} 5d^6$  osmium
- 15)  $[Rn] 7s^2 5f^{11}$  einsteinium *note: no d's, so the p. table referenced has Ac as 5f*
- 16)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 4d^{10} 4p^5$  not valid (take a look at "4d"; must be 3 d not 4!)
- 17)  $1s^2 2s^2 2p^6 3s^3 3d^5$  not valid (3p comes after 3s); also no 3 electrons in s
- 18)  $[Ra] 7s^2 5f^8$  not valid (radium isn't a noble gas)
- 19)  $[Kr] 5s^2 4d^{10} 5p^3$  valid
- 20)  $[Xe]$  not valid (an element can't be its own electron configuration)

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**RJ Alexander**



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## **Table of Contents Electron Configuration Worksheet And Lots More Answers Key Quizlet**

1. Understanding the eBook Electron Configuration Worksheet And Lots More Answers Key Quizlet
  - The Rise of Digital Reading Electron Configuration Worksheet And Lots More Answers Key Quizlet
  - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Worksheet And Lots More Answers Key Quizlet
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Electron Configuration Worksheet And Lots More Answers Key Quizlet
  - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Worksheet And Lots More Answers Key Quizlet
  - Personalized Recommendations

- Electron Configuration Worksheet And Lots More Answers Key Quizlet User Reviews and Ratings
- Electron Configuration Worksheet And Lots More Answers Key Quizlet and Bestseller Lists
- 5. Accessing Electron Configuration Worksheet And Lots More Answers Key Quizlet Free and Paid eBooks
  - Electron Configuration Worksheet And Lots More Answers Key Quizlet Public Domain eBooks
  - Electron Configuration Worksheet And Lots More Answers Key Quizlet eBook Subscription Services
  - Electron Configuration Worksheet And Lots More Answers Key Quizlet Budget-Friendly Options
- 6. Navigating Electron Configuration Worksheet And Lots More Answers Key Quizlet eBook Formats
  - ePub, PDF, MOBI, and More
  - Electron Configuration Worksheet And Lots More Answers Key Quizlet Compatibility with Devices
  - Electron Configuration Worksheet And Lots More Answers Key Quizlet Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Electron Configuration Worksheet And Lots More Answers Key Quizlet
  - Highlighting and Note-Taking Electron Configuration Worksheet And Lots More Answers Key Quizlet
  - Interactive Elements Electron Configuration Worksheet And Lots More Answers Key Quizlet
- 8. Staying Engaged with Electron Configuration Worksheet And Lots More Answers Key Quizlet
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Electron Configuration Worksheet And Lots More Answers Key Quizlet
- 9. Balancing eBooks and Physical Books Electron Configuration Worksheet And Lots More Answers Key Quizlet
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Electron Configuration Worksheet And Lots More Answers Key Quizlet
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Electron Configuration Worksheet And Lots More Answers Key Quizlet
  - Setting Reading Goals Electron Configuration Worksheet And Lots More Answers Key Quizlet
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electron Configuration Worksheet And Lots More Answers Key Quizlet
  - Fact-Checking eBook Content of Electron Configuration Worksheet And Lots More Answers Key Quizlet

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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