

Worksheet: Electron Configurations

Name: _____
Period: _____

1. Which electron configuration represents an atom in an excited state?

- A) $1s^2 2s^2 2p^6 3p^3$
 B) $1s^2 2s^2 2p^6 3s^2 3p^3$
 C) $1s^2 2s^2 2p^6 3s^2 3p^2$
 D) $1s^2 2s^2 2p^6 3s^2$

2. All of the elements in Period 3 have a total of 2 electrons in the

- A) $2s$ sublevel
 B) $3s$ sublevel
 C) $2p$ sublevel
 D) $3p$ sublevel

3. Which atom in the ground state has three unpaired electrons in its outermost principal energy level?

- A) Li
 B) B
 C) N
 D) Ne

4. What is the total number of valence electrons in an atom with the electron configuration $1s^2 2s^2 2p^6 3s^2 3p^3$?

- A) 6
 B) 2
 C) 3
 D) 5

5. Which orbital notation represents an atom of beryllium in the ground state?

- A) $1s$ $2s$ $\overbrace{2p}$

 B) $1s$ $2s$ $\overbrace{2p}$

 C) $1s$ $2s$ $\overbrace{2p}$

 D) $1s$ $2s$ $\overbrace{2p}$

6. Which element has a completely filled third principal energy level?

- A) Ar
 B) N
 C) Fe
 D) Zn

7. Which atom in the ground state has only 3 electrons in the $3p$ sublevel?

- A) phosphorus
 B) potassium
 C) argon
 D) aluminum

8. What is the total number of occupied principal energy levels in a neutral atom of neon in the ground state?

- A) 1
 B) 2
 C) 3
 D) 4

9. Which sublevel configuration correctly represents a completely filled third principal energy level?

- A) $3s^2 3p^6 3d^6$
 B) $3s^2 3p^6 3d^{10}$
 C) $3s^2 2p^6 3d^{10}$
 D) $3s^2 3p^6 3d^5$

10. Which atom in the ground state contains a partially filled $3p$ orbital?

- A) argon
 B) calcium
 C) potassium
 D) aluminum

11. In the ground state, the atoms of elements in Period 2 all have the same number of

- A) protons
 B) neutrons
 C) $1s$ electrons
 D) oxidation states

12. Which is the electron configuration for a neutral atom in the ground state?

- A) $1s^2 2s^2 3s^1$
 B) $1s^2 2s^2 2p^6 3s^1$
 C) $1s^2 2s^2 2p^6 3p^1$
 D) $1s^2 2s^2 2p^6 3s^1$

13. Which represents the electron configuration of an isotope of oxygen in the ground state?

- A) $1s^2 2s^2 2p^1$
 B) $1s^2 2s^2 2p^2$
 C) $1s^2 2s^2 2p^3$
 D) $1s^2 2s^2 2p^4$

14. The total number of sublevels in the fourth principal energy level of an atom is

- A) 1
 B) 2
 C) 3
 D) 4

15. Which is a possible electron configuration for argon in the excited state?

- A) $1s^2 2s^2 2p^6 3s^2 3p^2$
 B) $1s^2 2s^2 2p^6 3s^2 3p^3$
 C) $1s^2 2s^2 2p^6 3s^1 3p^7$
 D) $1s^2 2s^2 2p^6 3s^2 3p^4 4s^1$

16. Which is a correct description of the shape and spatial orientation of the p orbitals in an atom?

- A) All have the same shape but a different orientation.
 B) All have the same shape and the same orientation.
 C) All have a different shape and a different orientation.
 D) All have a different shape but the same orientation.

17. A maximum of 6 electrons can occupy

- A) an s orbital
 B) an s sublevel
 C) a p orbital
 D) a p sublevel

18. How does the ground state electron configuration of the hydrogen atom differ from that of a ground state helium atom?

- A) Hydrogen has one electron in a higher energy level.
 B) Hydrogen has two electrons in a lower energy level.
 C) Hydrogen contains a half-filled orbital.
 D) Hydrogen contains a completely filled orbital.

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