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J Rink

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for the nitration $4\,01\,$ ml of nitric acid $4\,00\,$ ml of sulfuric acid and $4\,56\,$ ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only $4\,56\,$ ml of bromobenzene was used the amount of product predicted to be recovered was $8\,7533\,$ g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction eas particularly nitration of methyl benzoate go give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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then actual is your actual findings yield actual theoretical x 100 what are the possible side reactions of nitration of bromobenzene dinitration can occur at ortho para or at both ortho positions occurs in overheating meta product happens very rarely because bromine is an ortho para director

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the bromine substituent in bromobenzene is an ortho para director the para position is favored due to steric effects in the reaction that takes place in this lab a nitro group is substituted onto bromobenzene at the para position to form 1 bromo 4 nitrobenzene

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on the right the arenium cation has a nitro set with incredibly small 0 6 angstrom bond distances this will result in an

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mixture of sulfuric acid or nitric acid creates what is known as a nitro group

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