

Answer Key

PG-1: What is the coefficient of the hydroxide ion for the half reaction when the half-reaction is balanced in basic solution? (Note the hydroxide ion is added, not subtracted.)

$\text{BrO}^- (\text{aq}) \rightarrow \text{Br}^- (\text{aq})$ unbalanced

144. 1 (C) 145. 2 (B) 146. 4



PG-2: What is the balanced half reaction of the species in reacting to form nitrogen monoxide in acidic solution?

147. $\text{NO}_3^- (\text{aq}) + 4\text{H}^+ (\text{aq}) \rightarrow \text{NO} (\text{g}) + 2\text{H}_2\text{O} (\text{l})$

(C) $\text{NO}_3^- (\text{aq}) + 2\text{H}^+ \rightarrow \text{NO} (\text{g}) + \text{H}_2\text{O} (\text{l})$



148. $\text{NO}_3^- (\text{aq}) + 2\text{H}^+ \rightarrow \text{NO} (\text{g}) + \text{H}_2\text{O} (\text{l})$

149. $\text{NO}_3^- (\text{aq}) + 2\text{H}^+ \rightarrow \text{NO} (\text{g}) + \text{H}_2\text{O} (\text{l})$

PG-3: What is the coefficient for water when the reaction is balanced in acidic solution? (Note the coefficient for water is added, not subtracted.)

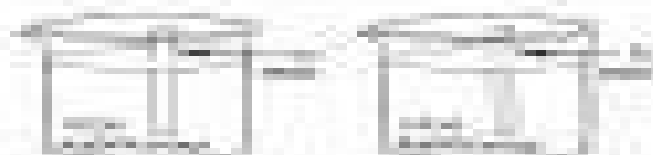


(C) 1 (B) 2 (A) 3 (D) 4



PG-4: Which reaction occurs in the cell in the presence of standard state conditions?

The standard reduction potential of zinc is $+1.26 \text{ V}$.
The standard reduction potential of iron is $+0.77 \text{ V}$.



(C) $\text{Zn} (\text{s}) \rightarrow \text{Zn}^{2+} (\text{aq})$

(B) $\text{Zn} (\text{s}) + \text{Fe}^{2+} (\text{aq}) \rightarrow \text{Zn}^{2+} (\text{aq}) + \text{Fe} (\text{s})$

(A) $\text{Zn}^{2+} (\text{aq}) + \text{Fe} (\text{s}) \rightarrow \text{Zn} (\text{s}) + \text{Fe}^{2+} (\text{aq})$

(D) $\text{Zn}^{2+} (\text{aq}) + \text{Fe}^{2+} (\text{aq}) \rightarrow \text{Zn} (\text{s}) + \text{Fe} (\text{s})$

The cell is a galvanic cell. The cell potential is $+0.49 \text{ V}$. The cell reaction is $\text{Zn} (\text{s}) + \text{Fe}^{2+} (\text{aq}) \rightarrow \text{Zn}^{2+} (\text{aq}) + \text{Fe} (\text{s})$.

For a reaction, the cell potential and the standard cell potential are the same. The cell potential is $+0.49 \text{ V}$. The standard cell potential is $+0.49 \text{ V}$. The cell reaction is $\text{Zn} (\text{s}) + \text{Fe}^{2+} (\text{aq}) \rightarrow \text{Zn}^{2+} (\text{aq}) + \text{Fe} (\text{s})$.

Electrochemistry Answers Study Guide For Content Mastery

Prof. Dr. C. Lingaraj



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Overcoming Students' Misconceptions in Science Mageswary Karpudewan,Ahmad Nurulazam Md Zain,A.L. Chandrasegaran,2017-02-28 This book discusses the importance of identifying and addressing misconceptions for the successful teaching and learning of science across all levels of science education from elementary school to high school It suggests teaching approaches based on research data to address students common misconceptions Detailed descriptions of how these instructional approaches can be incorporated into teaching and learning science are also included The science education literature extensively documents the findings of studies about students misconceptions or alternative conceptions about various science concepts Furthermore some of the studies involve systematic approaches to not only creating but also implementing instructional programs to reduce the incidence of these misconceptions among high school science students These studies however are largely unavailable to classroom practitioners partly because they are usually found in various science education journals that teachers have no time to refer to or are not readily available to them In response this book offers an essential and easily accessible guide International Society for Microbial Electrochemistry and Technology: Outputs From the 2018 Regional Meetings Sarah Glaven,Uwe Schröder,Eileen Hao Yu,Srikanth Mutnuri,Jeffrey A.

Gralnick,Chi Ho Chan,2020-07-09 **Ebook: Chemistry: The Molecular Nature of Matter and Change** Silberberg,2015-01-16 Ebook Chemistry The Molecular Nature of Matter and Change **My Chem Book (Class-12)** Prof. Dr. C. Lingaraj,2025-07-18 My Chem Book is a one stop solution for Class 12 Chemistry covering the CBSE NEET and JEE syllabi Based entirely on the NCERT framework it features concept based explanations previous year and NCERT questions and exam focused practice material The book is designed to strengthen core fundamentals and build strong application skills making Chemistry clear effective and approachable Even students without a strong base can study confidently as each chapter progresses step by step from basic concepts to advanced competitive level understanding The content is crafted from real classroom experience and addresses the actual difficulties students face Many students struggle in Chemistry not due to lack of effort but because of weak fundamentals which prevents them from understanding concepts and applying them in problem solving This often results in low scores My Chem Book is written with this in mind It simplifies each topic from the ground up and guides students in mastering both theory and application With the right study material and consistent involvement every student can succeed This book bridges the gap between textbook learning and true conceptual clarity By eliminating the need to refer to multiple books it saves time reduces confusion and acts as an exam saver for students By turning complex concepts into simple memorable ones My Chem Book empowers learners to study smart understand deeply and perform with confidence **Nanostructures** Osvaldo de Oliveira Jr,Marystela Ferreira,Alessandra Luzia Da Róz,Fabio de Lima Leite,2016-10-21 Nanostructures covers the main concepts and fundamentals of nanoscience emphasizing characteristics and properties of numerous nanostructures This book offers a clear explanation of nanostructured materials

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