

# REPORT SHEET

## Electrochemical Cells and Thermodynamics

EXPERIMENT

17

	Shorthand cell designation	Temperature (°C)	E cell (measured)	$\Delta G$ (calculated)	$K_{eq}$ (calculated)
1.	<del>Pb</del> Pb    Cu	20.2	.492 V	-94 kJ	1601.25
2.	Sn    Cu	<del>20.2</del> 20.3	.249 V	-78 kJ	2,198
3.	Sn    Pb	20.9	.125 V	-16 kJ	3.645

Show calculations for  $\Delta G^\circ$  and  $K_{eq}$  for an exemplary pair. For Pb | Pb<sup>2+</sup> || Sn<sup>2+</sup> | Sn:

Half-cell equation	E half-cell (calculated)
1. $(Pb(s) \rightarrow Pb^{2+} + 2e^-) (Cu^{2+} + 2e^- \rightarrow Cu)$	= .406 V
2. $(Sn \rightarrow Sn^{2+} + 2e^-) (Cu^{2+} + 2e^- \rightarrow Cu)$	= .425 V
3. $(Sn \rightarrow Sn^{2+} + 2e^-) (Pb^{2+} + 2e^- \rightarrow Pb)$	= .0196 V

### Effect of Temperature on Cell Potential

Cell designation: E (measured)	Temperature (°C)	Temperature (K)	$\Delta G$ (calculated)
Pb    Cu: .08 V	25	298	-16,598
Sn    Cu: .085 V	35	308	-16,405
Su    Pb: .086 V	50	323	-16,212
Pb    Cu: .082 V	60	333	-18,120
Su    Cu: .081 V	65	338	-18,653
Su    Pb: .079 V	70	343	-18,297

$\Delta S$  determined from the slope of a plot of  $\Delta G$  versus T 3.62

$\Delta F^\circ$  calculated at 298 K -7.652 (show calculations)

# Electrochemical Cells And Thermodynamics Lab Answers

**C Cleary**



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