

Unit 7 Worksheet 1: Molar-mass calculations

Four marks will be given for each question.

1. Write down balanced chemical equations

a) for the reaction of propanoic acid, acetic acid and acetic acid with calcium hydroxide

b) for the reaction of calcium hydroxide with calcium hydroxide

c) Please call up hydrogen in (b)

2. Hydrogen sulfide gas, which smells like rotten eggs, burns in air to produce water and sulfur. How many moles of hydrogen gas would be needed to completely burn 10 moles of hydrogen sulfide?

Equation: _____

Answer: _____

Change: _____

Notes: _____

3. Propane, C_3H_8 , burns in air to form carbon dioxide and water. If 10 moles of propane burn, how many moles of oxygen are needed?

Equation: _____

Answer: _____

Change: _____

Notes: _____

4. Acetylene, C_2H_2 , is used in the welding of metals by reacting hydrogen gas with oxygen gas to form hydrogen gas and acetylene. How many moles of acetylene gas are needed to burn 10 moles of hydrogen gas?

Equation: _____

Answer: _____

Change: _____

Notes: _____

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Ying Liu



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