

- 12 A sample of propanone,  $\text{CH}_3\text{COCH}_3$ , was heated under reflux with potassium dichromate(VI) acidified with sulfuric acid, and then the mixture was distilled. Apart from the peaks due to the  $\text{C}-\text{C}$  and  $\text{C}-\text{H}$  bonds, what peak(s) would be present in the IR spectrum of the distillate?

- ☐ A A peak due to the  $\text{C}=\text{O}$  only.  
☐ B A peak due to the  $\text{O}-\text{H}$  only.  
☐ C Peaks due to  $\text{C}=\text{O}$  and  $\text{O}-\text{H}$ .  
☐ D Peaks due to  $\text{C}-\text{O}$ ,  $\text{C}=\text{O}$  and  $\text{O}-\text{H}$ .

(Total for Question 12 = 1 mark)

- 13 Which of the following describes the appearance of iodine under the stated conditions?

	Solid	Dissolved in aqueous potassium iodide	Dissolved in a liquid hydrocarbon
<input type="checkbox"/> A	purple	brown	purple
<input type="checkbox"/> B	brown	blue-black	yellow
<input type="checkbox"/> C	shiny grey	brown	purple
<input type="checkbox"/> D	shiny grey	brown	brown

(Total for Question 13 = 1 mark)

- 14 Why do calculations of global warming exclude the effect of water vapour in the atmosphere?

- ☐ A Water vapour is not a greenhouse gas.  
☐ B Water vapour is a much less potent greenhouse gas than carbon dioxide.  
☐ C The average concentration of water vapour in the atmosphere is fairly constant.  
☐ D The concentration of water vapour in the atmosphere is much lower than that of carbon dioxide.

(Total for Question 14 = 1 mark)

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**RD Boyd**



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