

Limiting Reactant Worksheet #2

1. Given the following reaction: (Balance the equation first!)



- If you start with 14.8 g of C_3H_8 and 3.44 g of O_2 , determine the limiting reactant.
- Determine the number of moles of carbon dioxide produced.
- Determine the number of grams of H_2O produced.
- Determine the number of grams of excess reactant left.

2. Given the following equation:



- If 10.0 g of $\text{Al}_2(\text{SO}_4)_3$ is reacted with 10.0 g of NaOH , determine the limiting reactant.
- Determine the number of moles of $\text{Al}(\text{OH})_3$ produced.
- Determine the number of grams of Na_2SO_4 produced.
- Determine the number of grams of excess reactant left over in the reaction.

3. Given the following equation:



- If 25.4 g of Al_2O_3 is reacted with 10.2 g of Fe , determine the limiting reactant.
- Determine the number of moles of Al produced.
- Determine the number of grams of Fe_3O_4 produced.
- Determine the number of grams of excess reactant left over in the reaction.

4. Consider the reaction



- 80.0 grams of I_2O_5 , reacts with 28.0 grams of carbon monoxide, CO . Determine the mass of iodine I_2 , which could be produced?
- If, in the above situation, only 0.160 moles, of iodine, I_2 was produced.
 - What mass of iodine was produced?
 - What percentage yield of iodine was produced?

5. Zinc and sulfur react to form zinc sulfide according to the equation



If 25.0 g of zinc and 30.0 g of sulfur are mixed,

- Which chemical is the limiting reactant?
- How many grams of ZnS will be formed?
- How many grams of the excess reactant will remain after the reaction is over?

6. Which element is in excess when 3.00 grams of Mg is ignited in 2.20 grams of pure oxygen? What mass is in excess? What mass of MgO is formed?

7. How many grams of Al_2S_3 are formed when 5.00 grams of Al is heated with 10.0 grams S ?

Limiting Reagent Answer Key

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Table of Contents Limiting Reagent Answer Key

1. Understanding the eBook Limiting Reagent Answer Key
 - The Rise of Digital Reading Limiting Reagent Answer Key
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting Reagent Answer Key
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting Reagent Answer Key
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting Reagent Answer Key
 - Personalized Recommendations

- Limiting Reagent Answer Key User Reviews and Ratings
- Limiting Reagent Answer Key and Bestseller Lists
- 5. Accessing Limiting Reagent Answer Key Free and Paid eBooks
 - Limiting Reagent Answer Key Public Domain eBooks
 - Limiting Reagent Answer Key eBook Subscription Services
 - Limiting Reagent Answer Key Budget-Friendly Options
- 6. Navigating Limiting Reagent Answer Key eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting Reagent Answer Key Compatibility with Devices
 - Limiting Reagent Answer Key Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting Reagent Answer Key
 - Highlighting and Note-Taking Limiting Reagent Answer Key
 - Interactive Elements Limiting Reagent Answer Key
- 8. Staying Engaged with Limiting Reagent Answer Key
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting Reagent Answer Key
- 9. Balancing eBooks and Physical Books Limiting Reagent Answer Key
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting Reagent Answer Key
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Limiting Reagent Answer Key
 - Setting Reading Goals Limiting Reagent Answer Key
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Limiting Reagent Answer Key
 - Fact-Checking eBook Content of Limiting Reagent Answer Key

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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