

General Chemistry 1

Quarter 1, Lesson 14

LIMITING AND EXCESS REAGENTS

Objectives:

- Define the concept of limiting and excess reactants.
- Determine the limiting and excess reactants using mole ratios from balanced chemical equations.

When a chemical reaction is carried out in a flask, the amount of reactants are not always present in the exact proportion stoichiometrically determined from the balanced equation. To make sure that the more expensive reagent is completely used up and converted to the desired product, chemists usually add the cheaper reagent in excess quantity. The reagent that has completely reacted and used up in a reaction is called the **limiting reagent**. The **excess reagent** is the reactant that is present in quantity higher than what is required to react with the limiting reagent.

Sample Problem 1. Let's use the chemical equation below:



Ammonia, NH_3 , is synthesized from the reaction of H_2 and N_2 gases. Suppose 6 moles of H_2 was initially mixed with 4 moles of N_2 gas at high pressure. To determine which of the 2 reactants is the limiting reagent, the amount of NH_3 produced must be computed given the number of moles of H_2 and N_2 and the mole-mole factor from the balanced equation.

$$\text{moles } NH_3 = \cancel{\# \text{ moles of } H_2} \times \frac{2 \text{ moles } NH_3}{\cancel{3 \text{ moles } H_2}}$$

$$\begin{aligned} \text{moles } NH_3 &= 6 \cancel{\text{ moles of } H_2} \times \frac{2 \text{ moles } NH_3}{\cancel{3 \text{ moles } H_2}} \\ \text{moles } NH_3 &= 4 \text{ moles} \end{aligned}$$

$$\text{moles } NH_3 = \cancel{\text{moles of } N_2} \times \frac{2 \text{ moles } NH_3}{\cancel{1 \text{ mole } N_2}}$$

$$\begin{aligned} \text{moles } NH_3 &= 4 \cancel{\text{ moles of } N_2} \times \frac{2 \text{ moles } NH_3}{\cancel{1 \text{ mole } N_2}} \\ \text{moles } NH_3 &= 8 \text{ moles} \end{aligned}$$

From the computed values we determined that if 6 moles of H_2 completely reacts with N_2 , it can produce 4 moles of NH_3 while 4 moles of N_2 can produce 8 moles of NH_3 when fully used up. Since there is less amount of NH_3 produced with 6 moles of H_2 than 4 moles of N_2 , H_2 gas is the limiting reagent while the N_2 gas is the reagent in excess. To determine how much of the 4 moles of N_2 is in excess, we will use the mole-mole factor of N_2 and H_2 .

$$\begin{aligned} \text{moles } N_2 &= 6 \cancel{\text{ moles of } H_2} \times \frac{1 \text{ mole } N_2}{\cancel{3 \text{ moles } H_2}} \\ \text{moles } N_2 &= 2 \text{ moles} \end{aligned}$$

The number of moles of N_2 required to fully react to 6 moles of H_2 is only 2 moles. Thus, the initial 4 moles of N_2 has an excess of 2 moles.

$$\text{excess } N_2 = 2 \text{ moles} - 2 \text{ moles} = 2 \text{ moles}$$

Sample Problem 2. The combustion of ethane produces carbon dioxide and water shown in the reaction below:



1. How many moles of CO_2 are produced with 56.2 moles C_2H_6 and 73.4 moles of O_2 ?
Solution:

$$\begin{aligned} \text{moles } CO_2 &= 56.2 \cancel{\text{ moles of } C_2H_6} \times \frac{4 \text{ moles } CO_2}{\cancel{2 \text{ moles } C_2H_6}} \\ \text{moles } CO_2 &= 112.4 \text{ moles} \end{aligned}$$

$$\begin{aligned} \text{moles } CO_2 &= 73.4 \cancel{\text{ moles of } O_2} \times \frac{4 \text{ moles } CO_2}{\cancel{7 \text{ moles } O_2}} \\ \text{moles } CO_2 &= 41.9 \text{ moles} \end{aligned}$$

2. Which reagent is the limiting reagent? Which is the excess reagent?
The Limiting reagent is O_2 while the excess reagent is C_2H_6 .
3. How many moles of C_2H_6 are in excess?

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1. Understanding the eBook Limiting And Excess Reactants Answer Key 15
 - The Rise of Digital Reading Limiting And Excess Reactants Answer Key 15
 - Advantages of eBooks Over Traditional Books
2. Identifying Limiting And Excess Reactants Answer Key 15
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Limiting And Excess Reactants Answer Key 15
 - User-Friendly Interface
4. Exploring eBook Recommendations from Limiting And Excess Reactants Answer Key 15
 - Personalized Recommendations
 - Limiting And Excess Reactants Answer Key 15 User Reviews and Ratings
 - Limiting And Excess Reactants Answer Key 15 and Bestseller Lists
5. Accessing Limiting And Excess Reactants Answer Key 15 Free and Paid eBooks
 - Limiting And Excess Reactants Answer Key 15 Public Domain eBooks
 - Limiting And Excess Reactants Answer Key 15 eBook Subscription Services
 - Limiting And Excess Reactants Answer Key 15 Budget-Friendly Options

6. Navigating Limiting And Excess Reactants Answer Key 15 eBook Formats
 - ePub, PDF, MOBI, and More
 - Limiting And Excess Reactants Answer Key 15 Compatibility with Devices
 - Limiting And Excess Reactants Answer Key 15 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Limiting And Excess Reactants Answer Key 15
 - Highlighting and Note-Taking Limiting And Excess Reactants Answer Key 15
 - Interactive Elements Limiting And Excess Reactants Answer Key 15
8. Staying Engaged with Limiting And Excess Reactants Answer Key 15
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Limiting And Excess Reactants Answer Key 15
9. Balancing eBooks and Physical Books Limiting And Excess Reactants Answer Key 15
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Limiting And Excess Reactants Answer Key 15
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Limiting And Excess Reactants Answer Key 15
 - Setting Reading Goals Limiting And Excess Reactants Answer Key 15
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Limiting And Excess Reactants Answer Key 15
 - Fact-Checking eBook Content of Limiting And Excess Reactants Answer Key 15
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
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web oct 16 2023 to the upper right hand corner under the keyboard tab of the solidworks customize dialog there are buttons for print list and copy list copy

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web press a key or key combination if a shortcut key is already assigned a message informs you if you choose to use the shortcut for the new command it is removed from the old command letters are shown in uppercase when executing keyboard shortcuts caps

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