

4. I have 75 grams of Carbon Dioxide. How many molecules of Carbon Dioxide do I have? and how many atoms of Oxygen do I have?

5. Here, 75 grams of carbonic acid is breaking down into 53.2 grams of  $\text{CO}_2$  and 21.8 grams of  $\text{H}_2\text{O}$ . What is the limiting reactant?

- a.  $\text{CO}_2$
- b.  $\text{H}_2\text{O}$
- c.  $\text{HCO}_3$
- d. There is none

6. Define Limiting Reactant and Excess Reactant. Which one determines the amount of product that can be produced?

7. Balance the following chemical equation:



...and determine what is the limiting reactant if there is 40.3 moles of Calcium Oxide and 200 grams of water.

8. In the following reaction below (unbalanced) there are 45 grams of Aluminum and 54 grams of Oxygen. Find the following:



- a. Limiting Reactant
- b. Excess Reactant
- c. Amount of Excess Reactant Left

# Limiting Reactants Study Guide

**Y Pai**



## **Limiting Reactants Study Guide:**

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