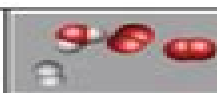


Activity A:**Water reaction**Get the Gizmo ready:

- Make sure $\text{H}_2 + \text{O}_2$ becomes H_2O is selected.
- Click **Reset** (↺).

**Goal: Identify a limiting reactant**

1. Count: H_2O is the **chemical formula** for water. In order to produce a single molecule of water, how many hydrogen atoms are needed? Oxygen atoms?

2. Predict: Set the number of O_2 molecules to five and the number of H_2 molecules to eight.

A. How many oxygen atoms are present? Hydrogen atoms?

B. How many water molecules could form from these reactants?

C. After the molecules react, which reactant will be left over?

D. Which reactant will be the limiting reactant?

E. Click **Play** and wait until **Reaction complete** is shown. What happened?

Due to the hydrogen and oxygen reacting, there are eight water molecules and one remaining oxygen molecule.

3. Test: Click **Reset**. Set the number of O_2 molecules to 10.

A. How many hydrogen molecules (H_2) will be needed for there to be no excess

reactants?
20
hydrogen
molecules

Use the slider to set the number of H_2 molecules to this value.

B. How many water molecules will be formed?

20

C. Click **Play**. What happened?

20 water molecules were formulated

4. Challenge yourself: Run the Gizmo with five O_2 molecules and nine H_2 molecules.

A. How many of each molecule do you think will be produced by this reaction?

O_2 :

H_2 :

H_2O :

B. Click **Play**. How many were actually produced? O_2 :

H_2 :

H_2O :

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Lifelique



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Limiting Reactant and Percent Lifeliqe,2019 This lesson plan covers analyzing chemical reactions in order to determine limiting reactants and excess reactants including calculating the amount of excess reactant calculate the theoretical yield of a reaction and calculating the percent yield of a reaction

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