Name: Date:

# **Limiting Reagent Worksheet**

Answers

Given the following reaction:

$$C_1H_A + \underline{5} O_1 \rightarrow \underline{3} CO_1 + \underline{4} H_1O_2$$

Balance the reaction.

Suppose 14.8 g of C<sub>3</sub>H<sub>8</sub> reacts with 3.44 g of O<sub>2</sub>. Determine the limiting reagent.

$$14.8 \text{ g C}_{3}H_{8} \times \frac{1 \text{ mol C}_{3}H_{8}}{44.09 \text{ g C}_{3}H_{8}} \times \frac{4 \text{ mol H}_{2}O}{1 \text{ mol C}_{3}H_{8}} \times \frac{18.02 \text{ g H}_{2}O}{1 \text{ mol H}_{2}O} = 24.2 \text{ g H}_{2}O$$

$$3.44 \text{ g O}_{2} \times \frac{1 \text{ mol O}_{2}}{32 \text{ g O}_{2}} \times \frac{4 \text{ mol H}_{2}O}{5 \text{ mol O}_{2}} \times \frac{18.02 \text{ g H}_{2}O}{1 \text{ mol H}_{2}O} = 1.55 \text{ g H}_{2}O$$

Since O, produces less quantity of H,O than C,H, does, it is the limiting reagent.

(b) Determine the amount of CO, produced.

Given the following reaction:

Balance the reaction.

What is the limiting reagent if 2.2 g of Mg reacts with 4.5 L of O, at STP?

2.2 g Mg \* 
$$\frac{1 \text{ mol Mg}}{24.3 \text{ g Mg}}$$
 \*  $\frac{2 \text{ mol MgO}}{2 \text{ mol Mg}}$  = 0.0905 mol MgO  
4.5 L O<sub>2</sub> \*  $\frac{1 \text{ mol O}_2}{22.4 \text{ L O}_2}$  \*  $\frac{2 \text{ mol MgO}}{1 \text{ mol O}_2}$  \* = 0.401 mol MgO

Since Mg produces less MgO than O, Mg is the limiting reagent.

Determine the amount of MgO produced in grams.

# **Limiting Reactant Work Sheet**

**M** Tight



Unveiling the Magic of Words: A Overview of "Limiting Reactant Work Sheet"

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