

Limiting Reagent Worksheet #1

Key

1. Given the following reaction: (Balance the equation first!)



- a. If you start with 14.8 g of C_3H_8 and 3.44 g of O_2 , determine the limiting reagent.

$$\frac{14.8\text{g C}_3\text{H}_8}{44.09\text{g C}_3\text{H}_8} \times \frac{1\text{mol C}_3\text{H}_8}{1\text{mol C}_3\text{H}_8} \times \frac{4\text{mol H}_2\text{O}}{1\text{mol C}_3\text{H}_8} \times \frac{18.02\text{g H}_2\text{O}}{1\text{mol H}_2\text{O}} = 24.2\text{g H}_2\text{O}$$

$$\frac{3.44\text{g O}_2}{32.00\text{g O}_2} \times \frac{1\text{mol O}_2}{1\text{mol O}_2} \times \frac{4\text{mol H}_2\text{O}}{5\text{mol O}_2} \times \frac{18.02\text{g H}_2\text{O}}{1\text{mol H}_2\text{O}} = 1.55\text{g H}_2\text{O}$$

$\therefore \text{O}_2$ is limiting

- b. Determine the number of grams of carbon dioxide produced.

$$\frac{3.44\text{g O}_2}{32.00\text{g O}_2} \times \frac{1\text{mol O}_2}{1\text{mol O}_2} \times \frac{3\text{mol CO}_2}{5\text{mol O}_2} \times \frac{44.01\text{g CO}_2}{1\text{mol CO}_2} = \boxed{2.81\text{g CO}_2}$$

- c. Determine the number of grams of H_2O produced.

$$\boxed{1.55\text{g H}_2\text{O}} \quad (\text{from a.})$$

- d. Determine the number of grams of excess reagent left.

$$\frac{3.44\text{g O}_2}{32.00\text{g O}_2} \times \frac{1\text{mol O}_2}{1\text{mol O}_2} \times \frac{1\text{mol C}_3\text{H}_8}{5\text{mol O}_2} \times \frac{44.09\text{g C}_3\text{H}_8}{1\text{mol C}_3\text{H}_8} = 0.948\text{g C}_3\text{H}_8 \text{ used}$$

$$14.8\text{g C}_3\text{H}_8 - 0.948\text{g C}_3\text{H}_8 = \boxed{13.9\text{g C}_3\text{H}_8 \text{ remain}}$$

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