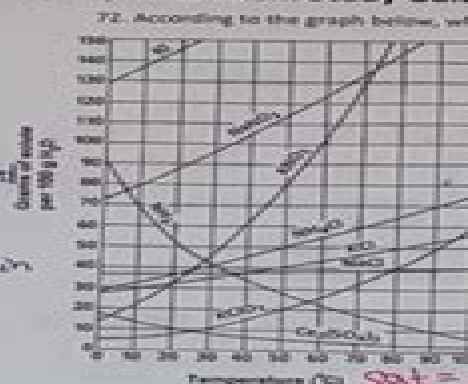


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64. Unsaturated
65. Saturated  
maximum amt. of solute dissolved  
at that temp.
66. Solubility  
the amount (g) of solute  
that can dissolve at a certain  
temp.
67. Rate of dissolution  
how fast a solute can  
dissolve
68. Molarity  
concentration  $M = \frac{\text{mol solute}}{1 \text{ Liter}}$
69. Supersaturated  
more than max amt. of solute  
dissolved in solvent
70. Electrolyte  
- must be soluble  
- must be ionic  
can conduct electricity
71. List 3 ways to increase the rate of dissolution of a solid:
1. increase surface area (crush)
  2. increase temperature
  3. stir

### SOLUBILITY OF COMMON IONIC COMPOUNDS IN WATER

<b>Insoluble compounds contain</b>	<b>Common exceptions</b>
$Cu_2(OH)_2CO_3$ , $CH_3COO^-$	None
$Fe^{3+}$	None
$FeO_4^{2-}$	None
$Cr^{3+}$	None
$CrO_4^{2-}$	None
$CrO_2^-$	None
$CrO_4^{2-}$	None
$Br^-$	Compounds of $Ag^+$ , $Hg_2^{2+}$ , and $Hg_2^{+}$
$Cl^-$	Compounds of $Ag^+$ , $Hg_2^{2+}$ , and $Hg_2^{+}$
$I^-$	Compounds of $Ag^+$ , $Hg_2^{2+}$ , and $Hg_2^{+}$
$SO_4^{2-}$	Compounds of $Sr^{2+}$ , $Ba^{2+}$ , $Pb^{2+}$ , and $Hg_2^{2+}$
<b>Insoluble compounds contain</b>	<b>Common exceptions</b>
$CO_3^{2-}$	Compounds of $Na^+$ and the alkali metal cations
$PO_4^{3-}$	Compounds of $Na^+$ and the alkali metal cations
$CrO_4^{2-}$	Compounds of $Na^+$ and the alkali metal cations
$Cr_2O_7^{2-}$	Compounds of $Na^+$ and the alkali metal cations
$OH^-$	Compounds of $Na^+$ , the alkali metal cations, $Ca^{2+}$ , $Sr^{2+}$ , and $Ba^{2+}$
$S^{2-}$	Compounds of $Na^+$ , the alkali metal cations, $Ca^{2+}$ , $Sr^{2+}$ , and $Ba^{2+}$



substance maintains almost constant solubility over the given temperature range?

nach

73. Which substance decreases in solubility with increasing temperature?

$$\text{NH}_3, \text{Ce}_2(\text{SO}_4)_3$$

74. What is the solubility of  $\text{KClO}_3$  at  $60^\circ\text{C}$ ?

$\sim 102.9$

75. A solution containing 50 g  $\text{CaCl}_2$  in 100 g  $\text{H}_2\text{O}$  is prepared at  $10^\circ\text{C}$ . Is the solution saturated, unsaturated or supersaturated?  $(100 \pm 50)$

(not on graph :))

76. Use the solubility rules table on the left to help you with the following:

	Soluble	Insoluble	Dissociation reaction	Is it an electrolyte?	How many ions are produced?
$\text{NaNO}_3$	✓		$\text{NaNO}_3 \xrightarrow{\text{aq}} \text{Na}^+ + \text{NO}_3^-$	✓	2
$\text{Na}_2\text{CO}_3$	✓		$\text{Na}_2\text{CO}_3 \xrightarrow{\text{aq}} 2\text{Na}^+ + \text{CO}_3^{2-}$	✓	3
$\text{NaCl}$	✓		$\text{NaCl} \xrightarrow{\text{aq}} \text{Na}^+ + \text{Cl}^-$	✓	2
$\text{Ca}(\text{NO}_3)_2$	✓		$\text{Ca}(\text{NO}_3)_2 \xrightarrow{\text{aq}} \text{Ca}^{2+} + 2\text{NO}_3^-$	✓	3
$\text{CaCO}_3$		✓	does not dissociate	NO	0

**Molarity (M)** is the number of moles of solute in one liter of solution

$$\text{molarity (M)} = \frac{\text{amount of solute (mol)}}{\text{volume of solution (L)}}$$

**Dilution:** the action of making a stock solution less concentrated

$$M_1V_1 = M_2V_2$$

77. Calculate the Molarity of a solution that contains 100grams of NaCl in 5.00L of solution( the molar mass of NaCl is 58.44 g/mol)

① convert g  $\rightarrow$  mol  

$$\frac{100 \text{ g NaCl}}{58.44 \text{ g}} \times \frac{1 \text{ mol}}{1} = 1.71 \text{ mol}$$

$$m = \frac{\text{mol}}{L}$$

$$M = \frac{1.71 \text{ mol}}{5.00 \text{ L}} =$$

0.34 molar solution NaCl

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