

Vocab:

64. Unsaturated

65. Saturated
maximum amt. of solute dissolved at that temp.

66. Solubility
the amount (g) of solute that can dissolve at a certain temp.

67. Rate of dissolution
how fast a solute can dissolve

68. Molarity
concentration $M = \frac{\text{mol solute}}{1 \text{ Liter}}$

69. Supersaturated
more than max amt. of solute dissolved in solvent

70. Electrolyte
- must be soluble
- must be ionic
can conduct electricity

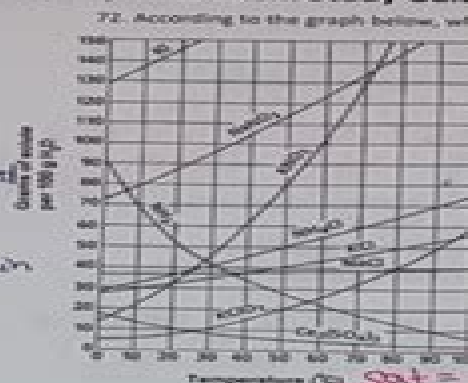
71. List 3 ways to increase the rate of dissolution of a solid:

1. increase surface area (crush)
2. increase temperature
3. stir

SOLUBILITY OF COMMON IONIC COMPOUNDS IN WATER

Soluble compounds contain	Common exceptions
$\text{Li}^+, \text{Na}^+, \text{K}^+, \text{NH}_4^+$	None
NO_3^-	None
Cl^-	None
ClO_4^-	None
ClO_3^-	None
ClO_2^-	None
Br^-	Compounds of Ag^+ , Pb^{2+} , and Hg_2^{2+}
I^-	Compounds of Ag^+ , Pb^{2+} , and Hg_2^{2+}
SO_4^{2-}	Compounds of Ag^+ , Pb^{2+} , and Hg_2^{2+}

Insoluble compounds contain	Common exceptions
CO_3^{2-}	Compounds of Na^+ and the alkali metal cations
PO_4^{3-}	Compounds of Na^+ and the alkali metal cations
CrO_4^{2-}	Compounds of Na^+ and the alkali metal cations
$\text{C}_2\text{O}_4^{2-}$	Compounds of Na^+ and the alkali metal cations
Br^-	Compounds of Ag^+ , Pb^{2+} , and Hg_2^{2+}
I^-	Compounds of Ag^+ , Pb^{2+} , and Hg_2^{2+}



72. According to the graph below, which substance maintains almost constant solubility over the given temperature range?

NaCl

73. Which substance decreases in solubility with increasing temperature?

 NH_3 , $\text{Ce}_2(\text{SO}_4)_3$

74. What is the solubility of KNO_3 at 60°C ?

 $\sim 102\text{g}$

75. A solution containing 50 g CaCl_2 in 100 g H_2O is prepared at 10°C . Is the solution saturated, unsaturated or supersaturated?

(Not on graph :)

76. Use the solubility rules table on the left to help you with the following:

	Soluble	Insoluble	Dissociation reaction	Is it an electrolyte?	How many ions are produced?
NaNO_3	✓		$\text{NaNO}_3 \rightarrow \text{Na}^+ + \text{NO}_3^-$	✓	2
Na_2CO_3	✓		$\text{Na}_2\text{CO}_3 \rightarrow 2\text{Na}^+ + \text{CO}_3^{2-}$	✓	3
NaCl	✓		$\text{NaCl} \rightarrow \text{Na}^+ + \text{Cl}^-$	✓	2
$\text{Ca}(\text{NO}_3)_2$	✓		$\text{Ca}(\text{NO}_3)_2 \rightarrow \text{Ca}^{2+} + 2\text{NO}_3^-$	✓	3
CaCO_3		✓	does not dissociate	NO	0

Molarity (M) is the number of moles of solute in one liter of solution

$$\text{molarity (M)} = \frac{\text{amount of solute (mol)}}{\text{volume of solution (L)}}$$

Dilution: the action of making a stock solution less concentrated

$$M_1V_1 = M_2V_2$$

77. Calculate the Molarity of a solution that contains 100 grams of NaCl in 5.00L of solution [the molar mass of NaCl is 58.44 g/mol]

$$\text{Convert g} \rightarrow \text{mol} \\ \frac{100 \text{ g NaCl}}{58.44 \text{ g/mol}} = 1.71 \text{ mol}$$

$$M = \frac{\text{mol}}{L}$$

$$M = \frac{1.71 \text{ mol}}{5.00 \text{ L}} =$$

0.34 molar solution NaCl

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A. A. Frempong



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