



Name:

Keith

Date:

Student Exploration: Boyle's Law and Charles's Law

Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: absolute zero, Boyle's law, Charles's law, Gay-Lussac's law, Kelvin scale, pressure

Prior Knowledge Question (Do this BEFORE using the Gizmo.)

A small helium tank measures about two feet (60 cm) high. Yet it can fill over 50 balloons! How can such a small tank contain enough helium to fill so many balloons?

the volume of the gas decreases and the tank can contain all the helium gas

Gizmo Warm-up

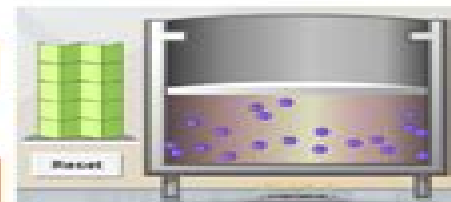
The Boyle's Law and Charles's Law Gizmo shows a container of gas. Inside, small purple spheres represent gas molecules.

1. Observe the particles. Are they all moving at the same speed?

No

2. How do the particles interact with the walls and lid of the container?

They bounce off of the lid and walls.



These interactions contribute to the **pressure** on the walls of the container. Pressure is defined as force per unit area. The SI units of pressure are newtons per square meter (N/m^2), or pascals (Pa).

2. Slowly drag the temperature (T) slider back and forth. (Note: In this Gizmo, the **Kelvin scale** is used to measure temperature. On the Kelvin scale, 0 degrees is **absolute zero**, the coldest possible temperature. Absolute zero is equal to -273.15°C or -459.67°F)

- A. How does the change in temperature affect the speed of the molecules?

It affects the pressure of the molecules.

- B. How does the change in temperature affect the volume of the container?

it affects how big or small the volume gets

Explore Learning Gizmo Answer Key Boyles Law

Sasha Hu



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