

Section 5.1: Models of the Atom

- How many sublevels are in the following principal energy levels? Which sublevels?
 - $n = 1$ **One - s**
 - $n = 2$ **Two - s, p**
 - $n = 3$ **Three - s, p, d**
 - $n = 4$ **Four - s, p, d, f**
 - $n = 5$ **Four - s, p, d, f**
 - $n = 6$ **Three - s, p, d**
 - $n = 7$ **Two - s, p**
- How many orbitals are in the following sublevels?
 - 1s sublevel **One**
 - 2p sublevel **Three**
 - 3d sublevel **Five**
 - 4f sublevel **Seven**
 - 5s sublevel **One**
 - 6p sublevel **Three**
- How many electrons can each of the following sublevels hold (maximum)?
 - s **Two**
 - p **Six**
 - d **Ten**
 - f **Fourteen**
- What is the shape of an s sublevel orbital? **Sphere**
- What is the shape of a p sublevel orbital? **Dumbbell**
- What was the problem with Rutherford's planetary model of the atom? **It could not explain chemical properties**
- Bohr proposed that electrons were found in specific circular paths called **orbits**
- What happens to the distance between energy levels as you increase in energy? **It decreases.**
- What does "electron cloud" mean? **A region of space with the highest probability, or likelihood, of finding an electron. Not an exact path.**
- How do electrons move from energy level to energy level? **Gain or lose energy**

Section 5.2: Electron Arrangement in Atoms

- Electrons arrange themselves around the nucleus in order to make the most **stable** arrangement.
- What are the three rules for writing electron configurations? Give the definition of each.
 - Aufbau's Principle** – electrons occupy the lowest energy levels first
 - Pauli Exclusion Principle** – electrons in the same orbital have opposite spin
 - Hund's Rule** – electrons enter orbitals individually before pairing up

Electrons In Atoms Chapter 5 Answer Key

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Electrons In Atoms Chapter 5 Answer Key:

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sentence second as a word equation and finally in its traditional symbolic form In this way students can focus on the meaning rather than the abstraction of the mathematics An appendix on English and SI units is also included Science by the Numbers To help students understand the importance of simple mathematical calculations in areas of magnitude several nontraditional calculations have been incorporated For example how much solid waste is generated in the United States how long it would take to erode a mountain and how many people were required to build Stonehenge Great Ideas and Great Ideas Concept Each chapter begins with a statement of a great unifying idea or theme in science and a concept map so that students immediately grasp the chief concept of the chapter and how the idea relates to the different branches of science These statements are intended to provide a framework for placing everyday experiences into a broad context Stop and Think Questions challenge students to think critically about the implications of a scientific discovery or principle Resources for Instructors and Students including practice quizzes flashcards lecture slides an instructor s manual images and tables from the book a test bank and much more

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