Electrons/Periodic Table Review Packet

Name

Period

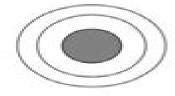
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Atomic Basics

Name ____

Part A: Atomic Structure

- 1. Draw five protons in the macleus of the atom. Label them with their charge.
- 2. Draw six neutrons in the nucleus of the atom.
- 3. Draw two electrons in the first energy level and label them with their charge.
- 4. Draw three electrons in the second energy level and label them with their charge.
- 5. What element is represented by the diagram?



Part B: Atomic Calculations

6. Label the information provided in the periodic table.

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gen •	
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- 7. What does the atomic number represent?
- 8. What does the atomic mass represent?
- 9. How would you figure the number of protons or electrons in an atom?
- 10. How would you figure the number of neutrons in an atom?
- 11. Use your knowledge of atomic calculations to complete the chart.

Element	Atomic Number	Atomic Mass	Protons	Neutrons	Electrons
Li	23	7			
P	243	258			
CI		28	27		
Ni	2963			2343	
34.	÷ :	200			100
Ag	407			Chil	
H		23	п		
Si	ij z			23:43	2343
W			76	ELED (D)	
Ne				23/80	23/00

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Electrons And The Periodic Table Review Answers

R Sandford

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