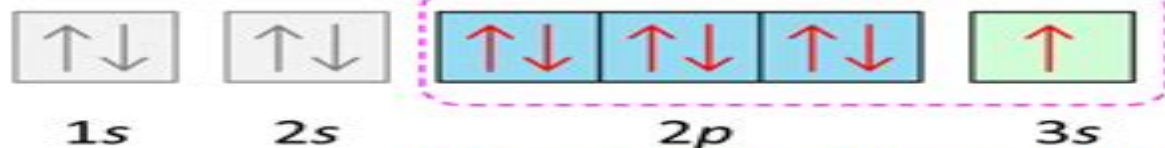


Electron Configuration of Main Group Cations

Electron configuration

Na $1s^2 2s^2 2p^6 3s^1$

Orbital diagram

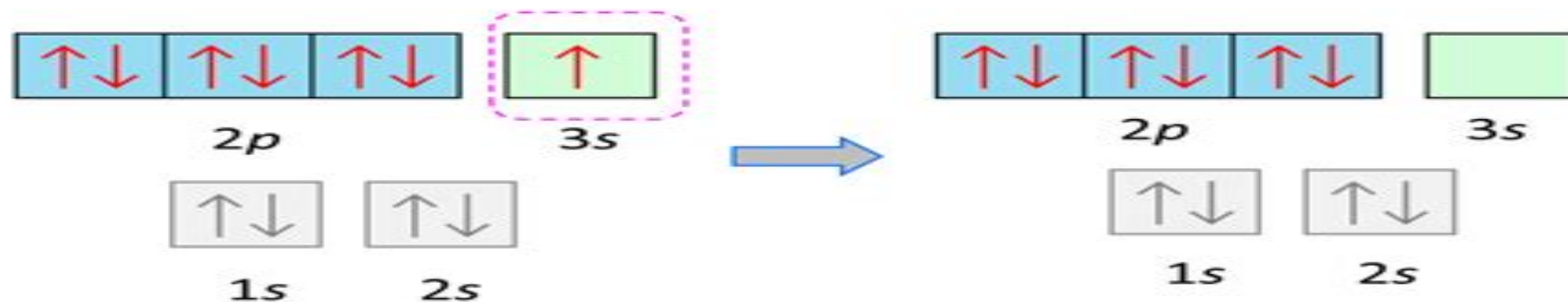


Focus on the valence shell

The charge is +1, therefore, one electron is lost:



The electron must be from the outermost valence orbital – $3s^1$



Na $1s^2 2s^2 2p^6 3s^1$

Na⁺ $1s^2 2s^2 2p^6$

↓ Same as

Ne $1s^2 2s^2 2p^6$



ChemistrySteps.com

Cations of main group metals attain the electron configuration of the noble gas that precedes them in the periodic table. They are **isoelectronic** with the previous noble gas.

Electron Configurations Of Atoms And Ions Lab Answers

Yan Bai



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Table of Contents Electron Configurations Of Atoms And Ions Lab Answers

1. Understanding the eBook Electron Configurations Of Atoms And Ions Lab Answers
 - The Rise of Digital Reading Electron Configurations Of Atoms And Ions Lab Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configurations Of Atoms And Ions Lab Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Electron Configurations Of Atoms And Ions Lab Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configurations Of Atoms And Ions Lab Answers
 - Personalized Recommendations
 - Electron Configurations Of Atoms And Ions Lab Answers User Reviews and Ratings
 - Electron Configurations Of Atoms And Ions Lab Answers and Bestseller Lists
5. Accessing Electron Configurations Of Atoms And Ions Lab Answers Free and Paid eBooks
 - Electron Configurations Of Atoms And Ions Lab Answers Public Domain eBooks
 - Electron Configurations Of Atoms And Ions Lab Answers eBook Subscription Services
 - Electron Configurations Of Atoms And Ions Lab Answers Budget-Friendly Options

6. Navigating Electron Configurations Of Atoms And Ions Lab Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Electron Configurations Of Atoms And Ions Lab Answers Compatibility with Devices
 - Electron Configurations Of Atoms And Ions Lab Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Electron Configurations Of Atoms And Ions Lab Answers
 - Highlighting and Note-Taking Electron Configurations Of Atoms And Ions Lab Answers
 - Interactive Elements Electron Configurations Of Atoms And Ions Lab Answers
8. Staying Engaged with Electron Configurations Of Atoms And Ions Lab Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Electron Configurations Of Atoms And Ions Lab Answers
9. Balancing eBooks and Physical Books Electron Configurations Of Atoms And Ions Lab Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Electron Configurations Of Atoms And Ions Lab Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Electron Configurations Of Atoms And Ions Lab Answers
 - Setting Reading Goals Electron Configurations Of Atoms And Ions Lab Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Electron Configurations Of Atoms And Ions Lab Answers
 - Fact-Checking eBook Content of Electron Configurations Of Atoms And Ions Lab Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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