

1. For **each** of the following sets of sublevels, determine which has the lowest energy, and which has the highest energy.

(a) 2s, 3s, 4s

(b) 4f, 4p, 4d

(c) 4d, 4f, 3d

(d) 6s, 3d, 5f

Ex for (a) - Lowest 2s, highest 4s

2. Explain what  $6s^2$  means.

3. What is the maximum number of electrons that can be found in an s – sublevel? \_\_\_\_ p – sublevel? \_\_\_\_ d – sublevel? \_\_\_\_ f – sublevel? \_\_\_\_

4. Indicate which of these (there may be more than one) orbital designations is incorrect:

(a) 4s

(b) 3f

(c) 2d

(d) 3d

5. Draw energy level diagrams for

(a) sodium

(b) silicon

(c) beryllium

(d) cobalt

6. How many unpaired electrons are there in each of the atoms in question 5?

7. Give the electron configuration for:

(a) lithium

(b) oxygen

(c) selenium

(d) iodine

(e) silver

(f) arsenic

(g) indium

(h) potassium

8. Give the name and symbol for the atom whose electron configuration corresponds to each of the following electron configurations.

a)  $1s^2 2s^2 2p^6 3s^2 3p^5$

b)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^3$

c)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^2 4d^6$

d)  $1s^2 2s^2 2p^6 3s^2 3p^2$

e)  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1$

9. Write the shorthand notation for the above electron configurations.

# Electron Configuration Exam Questions

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## **Table of Contents Electron Configuration Exam Questions**

1. Understanding the eBook Electron Configuration Exam Questions
  - The Rise of Digital Reading Electron Configuration Exam Questions
  - Advantages of eBooks Over Traditional Books
2. Identifying Electron Configuration Exam Questions
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Electron Configuration Exam Questions
  - User-Friendly Interface
4. Exploring eBook Recommendations from Electron Configuration Exam Questions

- Personalized Recommendations
- Electron Configuration Exam Questions User Reviews and Ratings
- Electron Configuration Exam Questions and Bestseller Lists
- 5. Accessing Electron Configuration Exam Questions Free and Paid eBooks
  - Electron Configuration Exam Questions Public Domain eBooks
  - Electron Configuration Exam Questions eBook Subscription Services
  - Electron Configuration Exam Questions Budget-Friendly Options
- 6. Navigating Electron Configuration Exam Questions eBook Formats
  - ePub, PDF, MOBI, and More
  - Electron Configuration Exam Questions Compatibility with Devices
  - Electron Configuration Exam Questions Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Electron Configuration Exam Questions
  - Highlighting and Note-Taking Electron Configuration Exam Questions
  - Interactive Elements Electron Configuration Exam Questions
- 8. Staying Engaged with Electron Configuration Exam Questions
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Electron Configuration Exam Questions
- 9. Balancing eBooks and Physical Books Electron Configuration Exam Questions
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Electron Configuration Exam Questions
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Electron Configuration Exam Questions
  - Setting Reading Goals Electron Configuration Exam Questions
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Electron Configuration Exam Questions

- Fact-Checking eBook Content of Electron Configuration Exam Questions
- Distinguishing Credible Sources

### 13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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