

8 • Electron Configurations & Periodicity

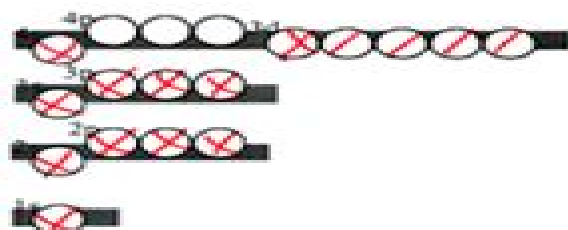
WRITING ELECTRON CONFIGURATIONS

For each given element, fill in the orbital diagram and then write the electron configuration for the element.

| | | | | | |
|---|---|---|--|--|---|
| 1. | 2. | 3. | 4. | 5. | 6. |
| | | | | | |
| Element: Ar # of e ⁻ s: <u>18</u> | Element: Mg # of e ⁻ s: <u>12</u> | Element: N # of e ⁻ s: <u>7</u> | Element: Li # of e ⁻ s: <u>3</u> | Element: P # of e ⁻ s: <u>15</u> | Element: Cl # of e ⁻ s: <u>17</u> |

Write the electron configurations of each of these in **long form** and **short form**:

- Ar $1s^2 2s^2 2p^6 3s^2 3p^6$
 Ar $[Ne] 3s^2 3p^6$
- Mg $1s^2 2s^2 2p^6 3s^2$
 Mg $[Ne] 3s^2$
- N $1s^2 2s^2 2p^3$
 N $[He] 2s^2 2p^3$
- Li $1s^2 2s^1$
 Li $[He] 2s^1$
- P $1s^2 2s^2 2p^6 3s^2 3p^3$
 P $[Ne] 3s^2 3p^3$
- Cl $1s^2 2s^2 2p^6 3s^2 3p^5$
 Cl $[Ne] 3s^2 3p^5$



- Fill in the orbital diagram for the element, Fe, and write the electron configuration of Fe in the long and short form.



Electron Configuration And Periodicity Lab Answers

Roman Wölfel



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